

Addition Compounds

Double Salt

Complex Salt

* Carnallite :-



* Mohr's Salt :-



* Potash Alum :-



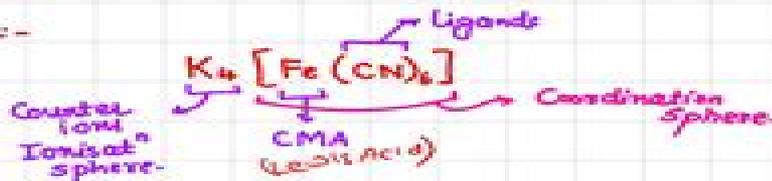
* d-block elements.

↳ * Small Size,

* high ϕ Nature.

* Vacant Orbitals

Note :-



* **Denticity** :- No. of e^- pair which can be donated by a ligand to metal ion.

* In any ligand if total no. of atoms are 3 or less than 3 \rightarrow then denticity is always "1".

↓
Stable
+ free

CLASSIFICATION OF LIGANDS :-

A) Anionic

SO_4^{2-} = Sulfate

CO_3^{2-} = Carbonate

Cl^- = Chloride

OH^- = Hydroxide

F^- = Fluoride

NH_2^- = Amide

NH^{2-} = Imide

N_3^- = Azide

N^3- = Nitride

O^{2-} = Oxide

O_2^- = Superoxide

O_2^{2-} = Peroxide

CN^- = Cyanide

NC^- = Isocyanide

NO_2^- = Nitrite-N

ONO^- = Nitrite-O

OCN^- = Cyanate-O

NCO^- = Cyanate-N

SCN^- = Thiocyanate-S

NCS^- = Thiocyanate-N

$C_2O_4^{2-}$ = Oxalate (2)

Gly = Glycinate (2)

dmg = dimethylglyoximate (2)

glyoximate (2)

$EDTA^{4-}$ = Ethylene diamine tetraacetate. (6)

B) Neutral :-

H_2O :- Aqua

NH_3 :- Amine

CO :- Carbonyl

NO :- Nitrosyl

PH_3 :- Phosphine

PPh_3 :- Triphenyl Phosphine

Py :- Pyridine

en :- ethylenediamine

pn :- propylene diamine

tn :- trimethylene diamine

dien :- diethylene triamine

trien :- triethylene tetraamine

C) Cationic :-

NO^+ = Nitrosonium

$NH_2-NH_3^+$:- Hydrazonium

Jugaad :-

$\rightarrow Pt/Pd$

always dsp^2

$CN=4$

O.S. = +2

$CN=6$

O.S. = +4

\rightarrow When Fe & NO both +1.

NO = +1

$\rightarrow Sc Ti V Cr Mn Fe Co Ni Cu Zn \} +2$

3d = 1 2 3 4 5 6 7 8 9 10

Similar for below 4d & 5d!

$\rightarrow d^1/d^2/d^3 \xrightarrow[6 \text{ ligands}]{sp^3d^2} d^2sp^3$

UPE :- 1 2 3 CN=6 always IOC

LSC Paramag.

$\rightarrow d^2/d^3/d^4 \xrightarrow[6 \text{ ligands}]{sp^3d^2} sp^3d^2$

UPE :- 2 1 0 CN=6 always OOC

HSC

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D Keegan



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