



14.2 Classifying Reactions

READ

Chemical reactions may be classified into different groups according to the reactants and products. The five major groups of chemical reactions are summarized below.

Synthesis reactions - when two or more substances combine to form a new compound.

- *General equation:* $A + B \rightarrow AB$
- *Example:* When rust forms, iron reacts with oxygen to form iron oxide (rust).
 $4\text{Fe (s)} + 3\text{O}_2 \text{(g)} \rightarrow 2\text{Fe}_2\text{O}_3 \text{(s)}$

Decomposition reactions - when a single compound is broken down to produce two or more smaller compounds.

- *General equation:* $AB \rightarrow A + B$
- *Example:* Water can be broken down into hydrogen and oxygen gases.
 $2\text{H}_2\text{O (l)} \rightarrow 2\text{H}_2 \text{(g)} + \text{O}_2 \text{(g)}$

Single displacement reactions - when one element replaces a similar element in a compound.

- *General equation:* $A + \text{BX} \rightarrow \text{AX} + \text{B}$
- *Example:* When iron is added to a solution of copper chloride, iron replaces copper in the solution and copper falls out of the solution.
 $\text{Fe (s)} + \text{CuCl}_2 \text{(aq)} \rightarrow \text{Cu (s)} + \text{FeCl}_2 \text{(aq)}$

Double displacement reactions - when ions from two compounds in solution exchange places to produce two new compounds.

- *General equation:* $\text{AX} + \text{BY} \rightarrow \text{AY} + \text{BX}$
- *Example:* When carbon dioxide gas is bubbled into lime water, a precipitate of calcium carbonate is formed along with water.
 $\text{CO}_2 \text{(g)} + \text{Ca(OH)}_2 \text{(aq)} \rightarrow \text{CaCO}_3 \text{(s)} + \text{H}_2\text{O (l)}$

Combustion reactions - when a carbon compound reacts with oxygen gas to produce carbon dioxide and water vapor. Energy is released from the reaction.

- *General equation:* $\text{Carbon Compound} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{energy}$
- *Example:* The combustion of methane gas.
 $\text{CH}_4 \text{(g)} + 2\text{O}_2 \rightarrow \text{CO}_2 \text{(g)} + 2\text{H}_2\text{O (g)}$

EXAMPLE

Classify the following reaction as synthesis, decomposition, single displacement, double displacement, or combustion. Explain your answer.



Answer: Displacement. Magnesium replaces copper in the compound.

Classifying Chemical Reactions Study Guide Answers

Marco Casella



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Whispering the Strategies of Language: An Emotional Quest through **Classifying Chemical Reactions Study Guide Answers**

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Table of Contents Classifying Chemical Reactions Study Guide Answers

1. Understanding the eBook Classifying Chemical Reactions Study Guide Answers
 - The Rise of Digital Reading Classifying Chemical Reactions Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Classifying Chemical Reactions Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Classifying Chemical Reactions Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Classifying Chemical Reactions Study Guide Answers

- Personalized Recommendations
 - Classifying Chemical Reactions Study Guide Answers User Reviews and Ratings
 - Classifying Chemical Reactions Study Guide Answers and Bestseller Lists
5. Accessing Classifying Chemical Reactions Study Guide Answers Free and Paid eBooks
- Classifying Chemical Reactions Study Guide Answers Public Domain eBooks
 - Classifying Chemical Reactions Study Guide Answers eBook Subscription Services
 - Classifying Chemical Reactions Study Guide Answers Budget-Friendly Options
6. Navigating Classifying Chemical Reactions Study Guide Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Classifying Chemical Reactions Study Guide Answers Compatibility with Devices
 - Classifying Chemical Reactions Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Classifying Chemical Reactions Study Guide Answers
 - Highlighting and Note-Taking Classifying Chemical Reactions Study Guide Answers
 - Interactive Elements Classifying Chemical Reactions Study Guide Answers
8. Staying Engaged with Classifying Chemical Reactions Study Guide Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Classifying Chemical Reactions Study Guide Answers
9. Balancing eBooks and Physical Books Classifying Chemical Reactions Study Guide Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Classifying Chemical Reactions Study Guide Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Classifying Chemical Reactions Study Guide Answers
- Setting Reading Goals Classifying Chemical Reactions Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Classifying Chemical Reactions Study Guide Answers

- Fact-Checking eBook Content of Classifying Chemical Reactions Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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