CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY

1) Which of the following are weak electrolytes? HCl, $HC_2H_3O_2$, NH_3, KCl

A) HCl, KCl
B) HCl, $HC_2H_3O_2$, NH_3, KCl
C) $HC_2H_3O_2$, KCl
D) $HC_2H_3O_2$, NH_3
E) HCl, $HC_2H_3O_2$, KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and HNO_3(aq)?

A) K^+ and H^+
B) H^+ and OH^-
C) K^+ and NO_3^-
D) H^+ and NO_3^-
E) OH^- only

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of $CaCO_3$ when aqueous solutions of Na_2CO_3 and $CaCl_2$ are mixed is _____.

A) $2Na^+(aq) + CO_3^{2-}(aq) \rightarrow Na_2CO_3(aq)$
B) $2Na^+(aq) + 2Cl^-(aq) \rightarrow 2NaCl(aq)$
C) $Na^+(aq) + Cl^-(aq) \rightarrow NaCl(aq)$
D) $Ca^{2+}(aq) + CO_3^{2-}(aq) \rightarrow CaCO_3(s)$
E) $Na_2CO_3(aq) + CaCl_2(aq) \rightarrow 2NaCl(aq) + CaCO_3(s)$

Answer: D Diff: 3

4) Which of the following are strong acids? HI, HNO_3, HF, HBr

A) HF, HBr
B) HI, HNO_3, HF, HBr
C) HI, HF, HBr
D) HNO_3, HF, HBr
E) HI, HNO_3, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces _____.

A) water and a salt
B) hydrogen gas
C) oxygen gas
D) sodium hydroxide
E) ammonia

Answer: A Diff: 2

6) Which of these metals is the <u>least</u> easily oxidized? Na, Au, Fe, Ca, Ag

A) Na
B) Au
C) Fe
D) Ca
E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?

$Pb(NO_3)_2(aq) + Ni(s) \rightarrow Ni(NO_3)_2(aq) + Pb(s)$
$Pb(NO_3)_2(aq) + Ag(s) \rightarrow$ No reaction
$Cu(NO_3)_2(aq) + Ni(s) \rightarrow Ni(NO_3)_2(aq) + Cu(s)$

A) Ni
B) Ag
C) Cu
D) Pb
E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100
B) 0.0500
C) 0.0333
D) 0.0250
E) 125

Answer: A Diff: 3

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