

Unit 3 Chem/Honors Chem Edgenuity Study Guide

The following topics will be covered on the unit test. You may not see **handwritten numbers** that your teacher will give you until you are a product of transformations under scrutiny from Edgenuity on this test.

- Periodic trends: atomic number, average atomic mass, radius
- Group use of the table
- Calculate and describe average atomic masses
- Periodic trends
 - o Atomic radius
 - o Ionic radius
 - o Ionization energy
 - o Electronegativity vs. electron affinity
- Electron configurations
 - o Longhand configuration
 - o Noble gas configuration
 - o Orbital notation
- Periodic Table History
 - o Dmitri Mendeleev
 - o Antoine Lavoisier
 - o Henry Cavendish
 - o Joseph Proust
 - o John Dalton
- Properties of metals, nonmetals, semi metals vs metalloids
- Periodic Table organization
 - o Groups vs periods
 - o Periodic table structure grouping of metalloids
 - o Noble gas vs other properties
- Quantum numbers
 - o n, l, m, s
- Electrons
- Bohr Model
- Bohr atomic model – calculate the energy levels
- Photoelectric effect and electron emission spectrum
- JJ Thomson's experiment
- Rutherford's experiment

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