Chemistry 1A Knick

Name: _____ **Your test is on Friday, 12/18/15**

Chemical Reactions Study Guide

Balancing Chemical Equations

The first step to balancing chemical reaction equations is to be able to accurately count the number of each type of element on both the reactants and products side of the arrow. This is done by multiplying the coefficient (big number in front of an element symbol) with the subscript (little number found after the element symbol). **Remember:** If no coefficient or subscript is written, then there is an implied, imaginary "1".

Examples: **3 Br_2** ← there are 6 bromine atoms ($3 \times 2 = 6$)
CaO ← there is 1 calcium and 1 oxygen (in each case $1 \times 1 = 1$) The "1"s are implied.

Practice: *How many of each type of atom is represented.*

3 Cl_2 __6__ **Fe** __1__ **2 S_8** __16__ **MgF_2** _____1 Mg:2F_____

I_2 __2__ **4 Ni** __4__ **2 Li_2O** ____4 Li: 2 O_____ **H_2O_2** _____2 H: 2O_____

H_2SO_4 _____2 H: 1 SO_4_____ **2 $Pb(NO_3)_2$** _____2 Pb: 4 NO_3_____

Now that we've learned to count, it's time to put our skills into action by altering the numbers of certain elements to suit our desires. In any case, our desires are to have the same number of each type of element on both the products and reactants side of the equation. We do this by **changing the coefficients**

Practice: *Balance each of the following chemical reaction equations.* **Check your answer** *by counting each element on each side when you think you're done.*

____2__H_2 + _____O_2 → ___2___H_2O

___16___Na + _____S_8 → __8____Na_2S

___2___$NaHCO_3$ → _____Na_2CO_3 + _____CO_2 + _____H_2O

____2__Al + ___3___CuO → _____Al_2O_3 + ___3___Cu

_____K_2O + _____H_2O → ___2___KOH

___6___Mg + _____P_4 → ____2__Mg_3P_2

_____C_2H_4 + ____3__O_2 → ___2___H_2O + ___2___CO_2

Chemistry Ws Answers Unit 4

Liying Dong

Chemistry Ws Answers Unit 4:

Unveiling the Energy of Verbal Beauty: An Mental Sojourn through **Chemistry Ws Answers Unit 4**

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