



Chapter Test Answers

CHAPTER 1: SOLUTIONS

Ans. 1. (a) 2. (d) 3. (a) 4. (c) 5. (a) 6. (b)

Ans. 7. (i) Acetone and aniline mixture represents negative deviations from Raoult's law, hence for this mixture, ΔH_{mix} and ΔV_{mix} is negative.

(ii) For positive deviations,

$$p_A > p_A^0 x_A \text{ and } p_B > p_B^0 x_B$$

(iii)



The above graph is correct for water and nitric acid mixture.

Water and nitric acid mixture shows negative deviations from Raoult's law, hence

$$p_A < p_A^0 x_A \text{ and } p_B < p_B^0 x_B$$

Or

Water-HCl mixture shows negative deviations from Raoult's law and solutions showing negative deviations from ideal behaviour form maximum boiling azeotrope.

Ans. 8. Gas B will have the higher value of K_{H_2} (Henry's Constant) as lower is the solubility of the gas in the liquid higher is the value of K_{H_2} .

Ans. 9. In case of liquids, due to increase of pressure inside the cooker, the boiling point of water increases leading to faster cooking than in pan.

Ans. 10.

Ideal solution	Non-ideal solution
The components of this solution obey Raoult's law at all temperatures and concentrations. i.e. $p_A = p_A^0 x_A$ and $p_B = p_B^0 x_B$	The components of this solution do not obey Raoult's law. They show positive and negative deviations from Raoult's law. i.e. $p_A \neq p_A^0 x_A$ and $p_B \neq p_B^0 x_B$
On mixing, there is no enthalpy and volume change. i.e. $\Delta H_{mix} = 0$ and $\Delta V_{mix} = 0$	On mixing, there is enthalpy and volume change. i.e. $\Delta H_{mix} \neq 0$ and $\Delta V_{mix} \neq 0$

Ans. 11.



TIP

(i) First find out ΔT_f by using the formula

$$\Delta T_f = \frac{K_f \times W_2 \times 1000}{M_2 \times W_1}$$

(ii) Then, find out the value of T_f by using formula.

$$\Delta T_f = T_f^0 - T_f$$

Given, W_2 (glucose) = 0.520 g

W_1 (H_2O) = 80.2 g

K_f (H_2O) = 1.86 Km^{-1}

M_2 of $C_6H_{12}O_6$ (glucose) = 180 g mol^{-1}

$$\Delta T_f = \frac{K_f \times W_2 \times 1000}{M_2 \times W_1}$$

$$= \frac{1.86 \text{ Km}^{-1} \times 0.520 \text{ g} \times 1000}{180 \text{ g mol}^{-1} \times 80.2 \text{ g}}$$

$$= 0.0669 \text{ K}$$

Freezing point of solution

$$T_f = 273 \text{ K} - 0.0669 \text{ K}$$

$$= 272.933 \text{ K}$$

Ans. 12. van't Hoff factor is the ratio of the experimental value of colligative property to the calculated value of the colligative property. It is used to find out the extent of dissociation or association.

In dissociation, observed molar mass has lesser value than normal, so i is more than one ($i > 1$), e.g. The value of i for ethanoic acid in benzene is nearly 0.5, while for aqueous KCl, NaCl and $MgSO_4$ solution, it is nearly 2. The value of i for K_2SO_4 is nearly 3.

Ans. 13. (i) $\Delta T_f = \frac{K_f W_2 \times 1000}{M_2 \times W_1}$

$$0.383 = \left(\frac{3.83 \times 2.56}{M \times 100} \right) \times 100$$

$$M = 256$$

$$5 \times x = 256$$

$$132 \times x = 256$$

$$x = 8$$

(ii) (a) If we place cells in solution containing 12% sodium chloride solution, they shrink.

(b) If we place blood cells in solution containing 0.4% sodium chloride solution, they swell.

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1. Understanding the eBook Chemistrymc Com Chapter Test Answers
 - The Rise of Digital Reading Chemistrymc Com Chapter Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistrymc Com Chapter Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistrymc Com Chapter Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistrymc Com Chapter Test Answers
 - Personalized Recommendations
 - Chemistrymc Com Chapter Test Answers User Reviews and Ratings
 - Chemistrymc Com Chapter Test Answers and Bestseller Lists
5. Accessing Chemistrymc Com Chapter Test Answers Free and Paid eBooks
 - Chemistrymc Com Chapter Test Answers Public Domain eBooks
 - Chemistrymc Com Chapter Test Answers eBook Subscription Services

- Chemistrymc Com Chapter Test Answers Budget-Friendly Options
- 6. Navigating Chemistrymc Com Chapter Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistrymc Com Chapter Test Answers Compatibility with Devices
 - Chemistrymc Com Chapter Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistrymc Com Chapter Test Answers
 - Highlighting and Note-Taking Chemistrymc Com Chapter Test Answers
 - Interactive Elements Chemistrymc Com Chapter Test Answers
- 8. Staying Engaged with Chemistrymc Com Chapter Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistrymc Com Chapter Test Answers
- 9. Balancing eBooks and Physical Books Chemistrymc Com Chapter Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistrymc Com Chapter Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistrymc Com Chapter Test Answers
 - Setting Reading Goals Chemistrymc Com Chapter Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistrymc Com Chapter Test Answers
 - Fact-Checking eBook Content of Chemistrymc Com Chapter Test Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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