

# Intro. to Equilibrium

## Introduction: Forward and Reverse Processes

Reversible processes are chemical reactions that proceed in both directions. For example, take the reaction of hydrogen gas and iodine gas to form hydrogen iodide:



If you start with only the reactants in this reaction, you will see a mixture of reactants and products. Over time, the system will reach a state of equilibrium. At this point, the forward and reverse processes are occurring at the same rate, and the concentrations of all species remain constant.

Reversible processes are also characterized by the fact that they can be reversed. In the case of the reaction above, if you start with only products, you will see a mixture of reactants and products. Over time, the system will reach a state of equilibrium. At this point, the forward and reverse processes are occurring at the same rate, and the concentrations of all species remain constant.



The law of mass action states that the equilibrium constant is given by:

$$K_c = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

The reaction quotient,  $Q$ , is a measure of the relative concentrations of the reactants and products at any point in time. If  $Q < K$ , the reaction will proceed in the forward direction. If  $Q > K$ , the reaction will proceed in the reverse direction. If  $Q = K$ , the system is at equilibrium.

## Critical Thinking Questions

- Consider the reversible reaction between hydrogen and iodine. In the beginning of the reaction, there are only reactants. As the reaction proceeds, the concentration of hydrogen iodide increases. What conditions must be met for the system to reach equilibrium? Describe the nature of equilibrium in terms of the forward and reverse processes. How do you know the system is at equilibrium? What happens if you change the concentration of one of the reactants or products? How do you know the system is at equilibrium? What happens if you change the temperature of the system? How do you know the system is at equilibrium? What happens if you change the volume of the system? How do you know the system is at equilibrium?

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