

# Intro. to Equilibrium

## Introduction: Forward and Reverse Processes

Reversible processes are chemical reactions that proceed in both directions. For example, take the reaction of hydrogen gas and iodine gas to form hydrogen iodide:



If you start with only the reactants, the reaction will proceed in the forward direction as the system begins to reach equilibrium. If you start with only the products, the reaction will proceed in the reverse direction as the system begins to reach equilibrium. At equilibrium, the forward and reverse processes occur at equal rates.

At equilibrium, the concentrations of reactants and products are constant. They are not zero, but they are constant. The system is in a state of dynamic equilibrium. The forward and reverse processes are occurring at equal rates.



The law of mass action states that at equilibrium:



The reaction quotient,  $Q_c$ , is a measure of the relative concentrations of reactants and products at any point in time. At equilibrium,  $Q_c = K_c$ . If  $Q_c < K_c$ , the reaction will proceed in the forward direction. If  $Q_c > K_c$ , the reaction will proceed in the reverse direction.

## Critical Thinking Questions

- Consider the reversible reaction between hydrogen and iodine. In the beginning of the reaction, there are only reactants. As the reaction proceeds, the number of hydrogen and iodine molecules decreases, and the number of hydrogen iodide molecules increases. When equilibrium is reached, the number of hydrogen and iodine molecules is constant, and the number of hydrogen iodide molecules is constant. What happens to the number of hydrogen and iodine molecules if you add more hydrogen gas to the system? What happens to the number of hydrogen iodide molecules? Explain your answer. (The number of hydrogen and iodine molecules will increase, and the number of hydrogen iodide molecules will decrease.)

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