

**AP[®] CHEMISTRY
2013 SCORING GUIDELINES**

**Question 1
(10 points)**

Answer the following questions about the solubility of some fluoride salts of alkaline earth metals.

- (a) A student prepares 100. mL of a saturated solution of MgF_2 by adding 0.50 g of solid MgF_2 to 100. mL of distilled water at 25°C and stirring until no more solid dissolves. (Assume that the volume of the undissolved MgF_2 is negligibly small.) The saturated solution is analyzed, and it is determined that $[\text{F}^-]$ in the solution is $2.4 \times 10^{-3} \text{ M}$.

- (i) Write the chemical equation for the dissolving of solid MgF_2 in water.

| | |
|---------------------------------------------------------------------------|---------------------------------------------|
| $\text{MgF}_2(s) \rightleftharpoons \text{Mg}^{2+}(aq) + 2\text{F}^-(aq)$ | 1 point is earned for the correct equation. |
|---------------------------------------------------------------------------|---------------------------------------------|

- (ii) Calculate the number of moles of MgF_2 that dissolved.

| | |
|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------|
| $\frac{2.4 \times 10^{-3} \text{ mol F}^-}{1.0 \text{ L}} \times 0.100 \text{ L} \times \frac{1 \text{ mol MgF}_2}{2 \text{ mol F}^-} = 1.2 \times 10^{-4} \text{ mol MgF}_2$ | 1 point is earned for the correct calculation of moles from concentration. 1 point is earned for the correct stoichiometry. |
|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------|

- (iii) Determine the value of the solubility-product constant, K_{sp} , for MgF_2 at 25°C .

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| $[\text{Mg}^{2+}] = \frac{1}{2} [\text{F}^-] = \frac{1}{2} (2.4 \times 10^{-3} \text{ M}) = 1.2 \times 10^{-3} \text{ M}$ $K_{sp} = [\text{Mg}^{2+}][\text{F}^-]^2 = (1.2 \times 10^{-3})(2.4 \times 10^{-3})^2$ $= 6.9 \times 10^{-9}$ | 1 point is earned for the correct value of $[\text{Mg}^{2+}]$ 1 point is earned for the correct setup for determining the value of K_{sp} . 1 point is earned for the correct value of K_{sp} . |
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Minjie Lin



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