

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH<sup>-</sup>] and [H<sup>+</sup>].
3. Calculate the pH of weak (or strong) acids.

## Vocabulary

- |                   |                                     |                   |
|-------------------|-------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair         | 3. weak acid      |
| 4. hydronium      | 5. hydroxide ion (OH <sup>-</sup> ) | 6. autoionization |
|                   | 7. acid solution                    | 8. pH             |

## Key Equations

- |   |                         |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$                             | 4. $pOH = -\log[OH^-]$  |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of 1.00. (Maximum possible score: 10.00)

1. What product is the \_\_\_\_\_ of a weak hydroxide ion and a hydronium ion? \_\_\_\_\_ of the conjugate acid-base pair.
2. \_\_\_\_\_ of a strong acid is the \_\_\_\_\_ of a weak acid.
3. \_\_\_\_\_ of a weak acid is the \_\_\_\_\_ of a strong acid.
4. \_\_\_\_\_ of a strong acid is the \_\_\_\_\_ of a weak acid.
5. \_\_\_\_\_ of a weak acid is the \_\_\_\_\_ of a strong acid.
6. \_\_\_\_\_ of a strong acid is the \_\_\_\_\_ of a weak acid.
7. \_\_\_\_\_ of a weak acid is the \_\_\_\_\_ of a strong acid.
8. \_\_\_\_\_ of a strong acid is the \_\_\_\_\_ of a weak acid.
9. \_\_\_\_\_ of a weak acid is the \_\_\_\_\_ of a strong acid.
10. \_\_\_\_\_ of a strong acid is the \_\_\_\_\_ of a weak acid.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**Rachel Sandford**



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