

**FIRST SEMESTER CHEMISTRY FINAL EXAMINATION FOR GRADE 12<sup>TH</sup>****CHOOSE THE CORRECT ANSWERS**

- If the solute-solvent interactions are greater than the solute-solute and solvent-solvent interactions, what will be the total vapor pressure of the solution?  
A. greater than that calculated from Raoult's law B. less than that calculated from Raoult's law  
C. The same as calculated from Raoult's law D. Raoult's law cannot be applied for such interaction
- What is the freezing point of a solution of 250g of  $\text{CaCl}_2$  in 1.0kg of water ( $K_f$  for water =  $-1.86^\circ\text{C}/\text{kg}/\text{mol}$ ) A.  $-1.3^\circ\text{C}$  B.  $-6.5^\circ\text{C}$  C.  $-9.0^\circ\text{C}$  D.  $-12.55^\circ\text{C}$
- What is the molarity of a 5 g of hydrogen peroxide in 100mL solution that is used for hair bleaching A. 0.015M B. 0.15M C. 1.5M D. 3M
- What is the osmotic pressure, in atm, of a 0.30M solution of glucose in water that is used for intravenous infusion at body temperature,  $37^\circ\text{C}$ ?  
A. 0.911 atm B. 773.2atm C. 7.6atm D. 76atm
- Given the unbalanced equation  
 $\text{KMnO}_4 + \text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{I}_2 + \text{H}_2\text{O}$  How many grams of  $\text{KMnO}_4$  are needed to make 500mL of 0.250N solution? A. 3.95g B. 1.49g C. 1.58g D. 63.2g
- In which of the following cases will the dissolution of sugar be the most rapid?  
A. Sugar crystals in hot water B. Sugar crystals in cold water  
C. Powdered sugar in hot water D. Powdered sugar in cold water
- Which of the following is not a colligative property  
A. Osmotic pressure B. Vapour pressure C. Boiling point D. Freezing point depression
- What is the normality of a 1M solution of  $\text{H}_3\text{PO}_4$  given the following reaction  
 $\text{H}_3\text{PO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{HPO}_4 + \text{H}_2\text{O}$  A. 3N B. 2N C. 6N D. 1N
- A given solution contains 100g of  $\text{NaCl}$  and 900g of water. what is the mole fraction of  $\text{NaCl}$   
A. 0.0331 B. 0.9669 C. 0.3 D. 0.7
- A 0.029m aqueous solution of  $\text{K}_2\text{SO}_4$  has a freezing point of  $-0.14^\circ\text{C}$ . what are the value of  $i$  factors ( $K_{f,\text{H}_2\text{O}} = -1.86^\circ\text{C}/\text{m}$ ) A. 2 B. 4 C. 2.3 D. 2.6
- Which of the following combination never form a buffer solution  
A.  $\text{HF}/\text{F}^-$  B.  $\text{HCOOH}/\text{HCOO}^-$  C.  $\text{HCO}_3^-/\text{CO}_3^{2-}$  D.  $\text{HNO}_3/\text{NO}_3^-$
- Which one of the following are true about the graph of weak acid and strong base  
A. The PH is equal to 7 at equivalence point B. The graph indicates the titration of strong base with strong acid  
C. The PH is greater than 7 at equivalence point D. at half neutralization  $\text{PH} = \text{PK}_a$
- A solution in which propionic acid is 0.94% ionized has a PH of 2.85. what is the values of the acid ionization constant ( $K_a$ ) for propionic acid A.  $4.2 \times 10^{-5}$  B.  $6.8 \times 10^{-4}$  C.  $1.3 \times 10^{-5}$  D.  $8.7 \times 10^{-7}$
- The acid-base indicator bromocresol green is a weak acid. The yellow acid and blue base forms of the indicator are present in equal concentrations in a solution when the PH is 4.68. what is the  $\text{PK}_a$  of bromocresol green A. 4.48 B. 4.68 C. 5.68 D. 6.68
- For the step wise dissociation of aqueous  $\text{H}_3\text{PO}_4$ , which of the following is not conjugate acid-base pair A.  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$  B.  $\text{H}_3\text{PO}_4$  and  $\text{H}_2\text{PO}_4^-$  C.  $\text{H}_2\text{PO}_4^-$  and  $\text{HPO}_4^{2-}$  D.  $\text{H}_2\text{PO}_4^-$  and  $\text{PO}_4^{3-}$
- To 0.2M solution of a weak monoprotic acid HA, enough quantity of its sodium salt NaA, was dissolved to give a concentration of 0.2M of the salt. what will be the acid concentration  $[\text{H}_3\text{O}^+]$  in the final solution ( $K_a$  of HA =  $1.8 \times 10^{-5}$ )  
A.  $1.8 \times 10^{-5}\text{M}$  B.  $1.8 \times 10^{-4}\text{M}$  C.  $2.8 \times 10^{-5}$  D.  $2.8 \times 10^{-4}$   
0.0057M B. 0.057M C. 0.57M D. 5.70M
- Consider the following equilibria I.  $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$  II.  $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NH}_3$ ,  
III.  $\text{HSO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{H}_2\text{SO}_3$  In which of the equilibrium does water acts as a Bronsted-lowry base?  
A. I only B. II only C. II and III D. I, II, III
- To a 50ml of a buffer solution that is 0.25M in HF and 0.1M in NaF, if 5ml of 0.05M NaOH is added, what is the final PH of this buffer solution,  $\text{PK}_a \text{ HF} = 3.17$  A. 2.77 B. 2.80 C. 2.74 D. 3.17
- When sodium acetate is added to an aqueous solution of acetic acid its PH  
A. increases B. Decreases C. Remain constant D. may increases or decreases
- How is a buffer solution prepared A. By mixing a weak acid and strong acid B. By mixing a weak acid and its conjugate base C. By mixing a strong acid and its conjugate base D. By mixing strong base and its conjugate acid
- Which one of the followings are not intensive properties A. color B. length C. density D. melting point

# 70 Question Chemistry Final Exam Answers

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**70 Question Chemistry Final Exam Answers:**

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**70 Question Chemistry Final Exam Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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