

Classifying Chemical Reactions

1. Balance each equation.
2. Classify the following reactions as synthesis, decomposition, combustion, single replacement, or double replacement.

	<u>Type of Reaction</u>
1. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$	decomposition
<hr/>	
2. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	double replacement
<hr/>	
3. $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$ (Propane)	combustion
<hr/>	
4. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$	single replacement
<hr/>	
5. $2\text{Al} + 3\text{NiBr}_2 \rightarrow 2\text{AlBr}_3 + 3\text{Ni}$	single replacement
<hr/>	
6. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$	synthesis & combustion
<hr/>	
7. $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ (Methane)	combustion
<hr/>	

Blank Classifying Chemical Reactions

Ian Pickup



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