

CHEMISTRY 9TH - Multiple Choice Questions

➔ CHEMICAL BONDING

CHAPTER# 04

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01. An example of the ionic compound is:
(a) H_2 (b) CH_4
(c) N_2 (d) \checkmark $NaCl$
02. Interaction between highly electron deficient hydrogen and highly electronegative atom is called
(a) \checkmark covalent bond (b) ionic bond
(c) hydrogen bond (d) metallic bond
03. Two fluorine atoms share one electron each in their outermost shell to achieve the electronic configuration of:
(a) Xe (b) Ar
(c) Kr (d) \checkmark Ne
04. Number of electrons lost by atoms of group IIIA equals:
(a) 1 (b) 2
(d) \checkmark 3 (d) 4
05. Atom which loses two electrons from its outer shell to form ion is called:
(a) oxygen (b) potassium
(c) \checkmark magnesium (d) carbon
06. In $NaCl$ crystal lattice each Na^+ ion is surrounded by:
(a) \checkmark 6 Cl^- ions (b) 6 Na^+ ions
(c) 8 Cl^- ions (d) 12 Cl^- ions
07. At room temperature most of ionic compounds are:
(a) amorphous solids (b) \checkmark crystal line solids
(c) Liquids (d) gases
08. Tendency of atoms to acquire eight electrons in their valence
(a) \checkmark octet rule (b) duplet rule
(c) triplet rule (d) none of above

Chapter 4 Florida Chemistry Answer

Annelies Wilder-Smith



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