

Chapter 5 Bonding Test A

Do Not Write on This Test!!! You may use a periodic table.

Multiple Choice-Identify the choice that best completes the statement or answers the question.

- Electrons involved in bonding between atoms are
 - valence electrons
 - inside the nucleus
 - closest to the nucleus
 - positively charged
- Each family in the periodic table has its own characteristic properties based on the number of
 - neutrons
 - valence electrons
 - protons
 - ions
- What is the greatest number of valence electrons an atom can have? With the exception of helium,
 - 2
 - 3
 - 8
 - 12
- If atoms of a halogen nonmetal (Group 17) gains one electron, the atoms then have
 - no valence electrons
 - 7 valence electrons
 - 8 valence electrons
 - 17 valence electrons
- When an atom loses an electron, it becomes a
 - positive ion
 - negative ion
 - neutral ion
 - neutral atom
- An ionic bond is the attraction between
 - similarly charged ions
 - oppositely charged ions
 - neutral ions
 - neutral atoms
- The element boron is directly above aluminum on the periodic table. Which statement about boron is true?
 - Boron is in the same period as aluminum
 - Boron is in the same group as aluminum and has 5 valence electrons
 - Boron has 5 valence electrons and is in the same period as aluminum
 - Boron is in the same group as aluminum and has 3 valence electrons
- Magnesium bromide is an ionic compound with the chemical formula $MgBr_2$. What does the "2" tell you?
 - Bromide has a 2- charge.
 - There are two magnesium ions to every bromide ion.
 - There are two bromide ions for every magnesium ion.
 - Bromide has a 2+ charge.
- What is the chemical name for the compound with the formula Na_2S ?
 - sodium fluoride
 - magnesium sulfide
 - lithium oxide
 - sodium sulfide
- In the chemical formula for an ionic compound, which item is written first?
 - positive ion
 - negative ion
 - subscript
 - charge
- Which of the following is a characteristic property of ionic compounds?
 - They have low melting points.

Chapter 5 Assessment For Chemistry

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