

## Electrons in Atoms Study Worksheet

Fill in the blanks in the following paragraph with the appropriate terms. In Bohr's model of the atom, electrons inhabit certain \_\_\_\_\_ levels, with the levels closest to the nucleus having \_\_\_\_\_ energy than those farther from the nucleus. In the \_\_\_\_\_ state of the atom, the electrons are in the lowest \_\_\_\_\_ level possible. When an atom absorbs energy, it is said to ascend to an \_\_\_\_\_ state, where it is unstable. The atom will soon \_\_\_\_\_ the same amount of energy absorbed which may be seen in the form of visible light. In the study of \_\_\_\_\_, this visible light is seen as the \_\_\_\_\_ spectrum of an element. This spectrum is \_\_\_\_\_ to each element, hence it is also called an element's "fingerprints". The modern view of light is that it has a \_\_\_\_\_ nature. In other words, light may behave as a stream of particles called \_\_\_\_\_ or \_\_\_\_\_, or light may behave as a \_\_\_\_\_.

In the wave view of light, the wave equation is often used to determine a wave's frequency or wavelength. The \_\_\_\_\_ is the distance between corresponding points on adjacent waves while the \_\_\_\_\_ is the number of waves passing a given point in a given time. The wave equation is: \_\_\_\_\_.

# Chapter 5 Electrons In Atoms Practice Problems Worksheet Answers

**H Kauffman**



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