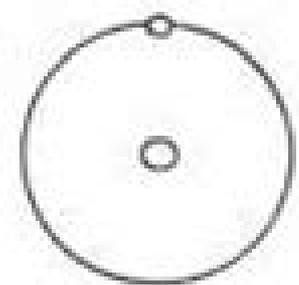


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Concept-Development Practice Page 32-1

Coulomb's Law

- 1. The diagram is of a hydrogen atom.
 - a. Label the proton in the nucleus with a + sign and the orbital electron with a - sign.
 - b. The electrical interaction between the nucleus and the orbital electron is a force of _____ (attraction) (repulsion).
 - c. According to Coulomb's law,



$$F = k \frac{q_1 q_2}{r^2}$$

If the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be

(greater) (less)

and if the distance between the nucleus and electron were greater, the force would be (greater) (less).

If the distance between the nucleus and electron were doubled, the force would be (1/4 as much) (1/2 as much) (two times as much) (4 times as much).

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Table of Contents Chapter 32 Conceptual Physics

1. Understanding the eBook Chapter 32 Conceptual Physics
 - The Rise of Digital Reading Chapter 32 Conceptual Physics
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 32 Conceptual Physics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 32 Conceptual Physics
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 32 Conceptual Physics
 - Personalized Recommendations
 - Chapter 32 Conceptual Physics User Reviews and Ratings
 - Chapter 32 Conceptual Physics and Bestseller Lists

5. Accessing Chapter 32 Conceptual Physics Free and Paid eBooks
 - Chapter 32 Conceptual Physics Public Domain eBooks
 - Chapter 32 Conceptual Physics eBook Subscription Services
 - Chapter 32 Conceptual Physics Budget-Friendly Options
6. Navigating Chapter 32 Conceptual Physics eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 32 Conceptual Physics Compatibility with Devices
 - Chapter 32 Conceptual Physics Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 32 Conceptual Physics
 - Highlighting and Note-Taking Chapter 32 Conceptual Physics
 - Interactive Elements Chapter 32 Conceptual Physics
8. Staying Engaged with Chapter 32 Conceptual Physics
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 32 Conceptual Physics
9. Balancing eBooks and Physical Books Chapter 32 Conceptual Physics
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 32 Conceptual Physics
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 32 Conceptual Physics
 - Setting Reading Goals Chapter 32 Conceptual Physics
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 32 Conceptual Physics
 - Fact-Checking eBook Content of Chapter 32 Conceptual Physics
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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