

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**
  - *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) = < 0.3)
    - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
  - *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
    - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
  - *Ionic Bond* ( $\Delta EN = > 1.7$ )
    - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
    - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
  - *Metallic Bond*
    - Bonding between two metallic atoms which occurs due to the attraction between a "sea of electrons" and the nuclei of bonded metal atoms
- \* Must memorize/know the difference in electronegativity ranges for each type of bond
  - \*\* Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
  - See the table at the bottom of Packet pg.5
  - You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
  - Why → because all atoms want to be at their LOWEST POTENTIAL ENERGY; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
  - How → ATTRACTIVE and REPULSIVE forces
    - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
    - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
  - **Bond Length**
    - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
  - **Bond Energy**
    - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chapter 6 Study Guide Chemistry

**VM Jensen**



## **Chapter 6 Study Guide Chemistry:**

The Top Books of the Year Chapter 6 Study Guide Chemistry The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous engrossing novels enthraling the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the fascinating narratives that have charmed audiences this year. Chapter 6 Study Guide Chemistry :

Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can triumph.

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Discover the Magic : Delia Owens "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting.

These popular novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of compelling stories waiting to be discovered.

The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

The Secret History is a brilliant and gripping novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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