

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (14.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____ of the autoionization of water which produces an acidic solution. _____
2. Hydrochloric acid ionizes from _____ to a strong acid. _____
3. _____ autoionization produces the acidic ion, _____ and the basic ion, _____.
4. _____ ionizes to form a weak acidic solution. _____
5. _____ ionizes to form a weak basic solution. _____
6. _____ autoionization produces _____ and _____.
7. _____ autoionization produces _____ and _____.

1. _____
2. _____
3. _____
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6. _____
7. _____
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9. _____
10. _____
11. _____
12. _____
13. _____
14. _____

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