

Chapter 7: Chemical Formulas and Chemical Compounds

I. CHEMICAL NAMES AND FORMULAS

A. Significance of a Chemical Formula

1. A chemical formula shows the relative number of atoms of each kind in a chemical compound, with subscripts indicating the number of atoms of one type of element.
2. In a molecular formula, the chemical formula shows the number of atoms of each element in one molecule of the compound.
 - a) For example, H_2O , the chemical formula for water, indicates the presence of 2 hydrogen atoms and 1 oxygen atom.
3. In an ionic compound, the chemical formula represents one formula unit, which is the simplest ratio of the compound, positive ions (cations) and negative ions (anions).

B. Monatomic Ions

1. Many main-group elements can become ions by gaining or losing electrons to have noble-gas configurations.
2. Monatomic ions are ions formed from a single atom.
 - a) Group 1 metals lose one electron; Group 2 elements lose two electrons. They will have +1 and +2 charges, respectively.
 - b) The nonmetals in Group 15, 16, and 17 gain electrons and form anions.
 - c) Metals with *d* orbitals are slightly different, as they can form +2, +3, or even +1 and +4. Some of them can even form ions of different charges.
 - i) Refer to the chart on page 221 of Modern Chemistry published by Holt, Rinehart, and Winston.

C. Naming Monatomic Ions

1. Monatomic cations are identified by their normal name.
2. Monatomic anions, however, have their ending dropped and the ending , -ide is added to the root name.

D. Binary Ionic Compounds

1. Binary compounds are compounds composed of two elements.
2. In a binary ionic compound, the total numbers of positive charges and negative charges must be equal.

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