

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_ a. ionic bond  
 \_\_\_\_\_ **bond energy** \_\_\_\_\_ b. covalent bond  
 \_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_ a. H—F \_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_ d. H—H  
 \_\_\_\_\_ **ionic** \_\_\_\_\_ b. Na—Cl \_\_\_\_\_ **polar covalent** \_\_\_\_\_ e. H—C  
 \_\_\_\_\_ **polar covalent** \_\_\_\_\_ c. H—O \_\_\_\_\_ **polar covalent** \_\_\_\_\_ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

# Chapter 8 Covalent Bonding Reading Guide Answer Key

**Emilie Sanchez**



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## **Table of Contents Chapter 8 Covalent Bonding Reading Guide Answer Key**

1. Understanding the eBook Chapter 8 Covalent Bonding Reading Guide Answer Key
  - The Rise of Digital Reading Chapter 8 Covalent Bonding Reading Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Reading Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Covalent Bonding Reading Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Reading Guide Answer Key

- Personalized Recommendations
  - Chapter 8 Covalent Bonding Reading Guide Answer Key User Reviews and Ratings
  - Chapter 8 Covalent Bonding Reading Guide Answer Key and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Reading Guide Answer Key Free and Paid eBooks
- Chapter 8 Covalent Bonding Reading Guide Answer Key Public Domain eBooks
  - Chapter 8 Covalent Bonding Reading Guide Answer Key eBook Subscription Services
  - Chapter 8 Covalent Bonding Reading Guide Answer Key Budget-Friendly Options
6. Navigating Chapter 8 Covalent Bonding Reading Guide Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Chapter 8 Covalent Bonding Reading Guide Answer Key Compatibility with Devices
  - Chapter 8 Covalent Bonding Reading Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Reading Guide Answer Key
  - Highlighting and Note-Taking Chapter 8 Covalent Bonding Reading Guide Answer Key
  - Interactive Elements Chapter 8 Covalent Bonding Reading Guide Answer Key
8. Staying Engaged with Chapter 8 Covalent Bonding Reading Guide Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 8 Covalent Bonding Reading Guide Answer Key
9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Reading Guide Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Reading Guide Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Reading Guide Answer Key
- Setting Reading Goals Chapter 8 Covalent Bonding Reading Guide Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Reading Guide Answer Key

- Fact-Checking eBook Content of Chapter 8 Covalent Bonding Reading Guide Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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