Name _____ Date _____ Class _____

CHAPTER 6 REVIEW

Chemical Bonding

MIXED REVIEW

SHORT ANSWER **Answer the following questions in the space provided.**

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

lattice energy	**a.** ionic bond
bond energy	**b.** covalent bond
enthalpy of vaporization	**c.** metallic bond

2. Use the electronegativity values shown in **Figure 20,** on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

ionic	**a.** H—F	**nonpolar covalent**	**d.** H—H
ionic	**b.** Na—Cl	**polar covalent**	**e.** H—C
polar covalent	**c.** H—O	**polar covalent**	**f.** H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H_2S and H_2O have similar structures and their central atoms belong to the same group. Yet H_2S is a gas at room temperature and H_2O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H_2O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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Chapter 8 Covalent Bonding Workbook Answer Key

Karin Nielsen-Saines

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