

Chapter Test B

Chapter: Chemical Equations and Reactions

Match a name space (possibly writing the name of the react or product molecule) completed with reactants/products atoms/molecules.

- The product(s) of a chemical reaction will be represented by a double arrow pointing towards the right. The substance(s) on the left side of the reaction is/are
 - reactant(s)**
 - product(s)
 - solvent
- An ionic chemical equation is represented by adding the
 - charges**
 - structure of the particles
 - number of particles
- An ionic equation shows the activity of aq ions. The other ions are
 - reactants
 - omitted as ions**
 - solvents
 - precipitates
- The activity for a precipitation reaction is shown as follows:
 - solvent molecules are added to the activity of the reactants
 - precipitates are formed
 - reactants are shown as ions
 - solvent molecules are added to the activity of the reactants**
- The coefficients are balanced equations
 - indicate the number of moles of each substance that react**
 - show the number of grams of each substance that react
 - show the number of molecules of the substances
 - show the relative volumes of each state
- If $2H_2 + O_2 \rightarrow 2H_2O$ then 2 moles of H_2 will react with 1 mole of O_2 to form
 - 2 moles of H_2O**
 - 1 mole of H_2O
 - 2 moles of H_2O
 - 1 mole of H_2O

Chapter 8 Test B

Tao Wei



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