

Chemistry: Chapter 9 Structure and properties of substances

Combined Science (Chemistry Part): Chapter 9 Structure and properties of substances

Multiple-choice questions

Sections 9.1–9.2

||CMEMF0020900001|| ||CSEMF0020900001||

Chlorine exists as a gas at room temperature and pressure because

- A. chlorine is reactive.
- B. the intermolecular forces among chlorine molecules are weak.
- C. the intermolecular forces between chlorine atoms are weak.
- D. chlorine molecules are diatomic.



##B##

||CMEMF0020900002|| ||CSEMF0020900002||

Helium has a low boiling point because

- A. a helium atom has a duplet structure.
- B. helium molecules are monoatomic.
- C. the van der Waals' forces between helium atoms are weak.
- D. helium is a noble gas.



##C##

||CMEMF0020900003|| ||CSEMF0020900003||

Which of the following statements concerning the structure of solid carbon dioxide is/are correct?

- (1) It has a simple molecular structure.
 - (2) Carbon atom and oxygen atom within the molecule are held by ionic bond.
 - (3) Carbon dioxide molecules are held together by intermolecular forces.
- A. (1) only
 - B. (2) only
 - C. (1) and (3) only
 - D. (2) and (3) only



##C Carbon dioxide is a covalent molecule. There are two C=O bonds joining the carbon and oxygen atoms together in the molecule.##

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Cary I. Sneider



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