

Science

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(Chapter - 1) (Chemical Reactions and Equations) (Intext)
(Class 10)

Page 10

Question 1:

A solution of a substance 'X' is used for white washing.

(i) Name the substance 'X' and write its formula.

(ii) Write the reaction of the substance 'X' named in (i) above with water.

Answer 1:

(i) The substance 'X' is calcium oxide. Its chemical formula is CaO.

(ii) Calcium oxide reacts vigorously with water to form calcium hydroxide (slaked lime)



Question 2:

Why is the amount of gas collected in one of the test tubes in Activity 1.7 double of the amount collected in the other? Name this gas.

Answer 2:

During the Electrolysis of water, hydrogen and oxygen is get separated by the electricity. Water (H₂O) contains two parts hydrogen and one part oxygen. Since hydrogen goes to one test tube and oxygen goes to another, the amount of gas collected in one of the test tubes is double of the amount collected in the other.

Page 13

Question 1:

Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

Answer 1:

When an iron nail is dipped in a copper sulphate solution, iron (which is more reactive than copper) displaces copper from copper sulphate solution forming iron sulphate, which is green in colour.



Therefore, the blue colour of copper sulphate solution fades and green colour appears.

Question 2:

Give an example of a double displacement reaction other than the one given in Activity 1.10.

Answer 2:

Sodium carbonate reacts with calcium chloride to form calcium carbonate and sodium chloride.



In this reaction, sodium carbonate and calcium chloride exchange ions to form two new compounds. Hence, it is a double displacement reaction.

Question 3:

Identify the substances that are oxidised and the substances that are reduced in the following reactions.

(i) $4\text{Na}(s) + \text{O}_2(g) \rightarrow 2\text{Na}_2\text{O}(s)$

(ii) $\text{CuO}(s) + \text{H}_2(g) \rightarrow \text{Cu}(s) + \text{H}_2\text{O}(l)$

Answer 3:

(i) Sodium (Na) is oxidised as it gains oxygen and oxygen gets reduced.

(ii) Copper oxide (CuO) is reduced to copper (Cu) while hydrogen (H₂) gets oxidised to water (H₂O).

Chapter Question Answers

M Mark



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Table of Contents Chapter Question Answers

1. Understanding the eBook Chapter Question Answers
 - The Rise of Digital Reading Chapter Question Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Question Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Question Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Question Answers
 - Personalized Recommendations
 - Chapter Question Answers User Reviews and Ratings
 - Chapter Question Answers and Bestseller Lists
5. Accessing Chapter Question Answers Free and Paid eBooks
 - Chapter Question Answers Public Domain eBooks
 - Chapter Question Answers eBook Subscription Services

- Chapter Question Answers Budget-Friendly Options
- 6. Navigating Chapter Question Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Question Answers Compatibility with Devices
 - Chapter Question Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Question Answers
 - Highlighting and Note-Taking Chapter Question Answers
 - Interactive Elements Chapter Question Answers
- 8. Staying Engaged with Chapter Question Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Question Answers
- 9. Balancing eBooks and Physical Books Chapter Question Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Question Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Question Answers
 - Setting Reading Goals Chapter Question Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Question Answers
 - Fact-Checking eBook Content of Chapter Question Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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