

# Chemistry

(www.tiwariacademy.com)  
(Chapter 1)(Solutions) (Exercises)  
(Class 12)

## Question 1.26:

If the density of some lake water is  $1.25 \text{ g mL}^{-1}$  and contains 92 g of  $\text{Na}^+$  ions per kg of water, calculate the molality of  $\text{Na}^+$  ions in the lake.

### Answer 1.26:

Number of moles present in 92 g of  $\text{Na}^+$  ions =  $\frac{92 \text{ g}}{23 \text{ g mol}^{-1}} = 4 \text{ mol}$

Therefore, molality of  $\text{Na}^+$  ions in the lake =  $\frac{4 \text{ mol}}{1 \text{ kg}} = 4 \text{ m}$

## Question 1.27:

If the solubility product of  $\text{CuS}$  is  $6 \times 10^{-16}$ , calculate the maximum molarity of  $\text{CuS}$  in aqueous solution.

### Answer 1.27:

Solubility product of  $\text{CuS}$ ,  $K_{sp} = 6 \times 10^{-16}$  Let  $s$  be the solubility of  $\text{CuS}$  in  $\text{mol L}^{-1}$ .



$$K_{sp} = [\text{Cu}^{2+}][\text{S}^{2-}]$$

Now,

$$= s \times s = s^2$$

Then, we have,  $K_{sp} = s^2 = 6 \times 10^{-16}$

$$\Rightarrow s = \sqrt{6 \times 10^{-16}} = 2.45 \times 10^{-8} \text{ mol L}^{-1}$$

Hence, the maximum molarity of  $\text{CuS}$  in an aqueous solution is  $2.45 \times 10^{-8} \text{ mol L}^{-1}$ .

## Question 1.28:

Calculate the mass percentage of aspirin ( $\text{C}_9\text{H}_8\text{O}_4$ ) in acetonitrile ( $\text{CH}_3\text{CN}$ ) when 6.5 g of  $\text{C}_9\text{H}_8\text{O}_4$  is dissolved in 450 g of  $\text{CH}_3\text{CN}$ .

### Answer 1.28:

6.5 g of  $\text{C}_9\text{H}_8\text{O}_4$  is dissolved in 450 g of  $\text{CH}_3\text{CN}$ .

Then, total mass of the solution =  $(6.5 + 450) \text{ g} = 456.5 \text{ g}$

Therefore, mass percentage of  $\text{C}_9\text{H}_8\text{O}_4 = \frac{6.5}{456.5} \times 100 \% = 1.424\%$

## Question 1.29:

Nalorphene ( $\text{C}_{19}\text{H}_{21}\text{NO}_3$ ), similar to morphine, is used to combat withdrawal symptoms in narcotic users. Dose of nalorphene generally given is 1.5 mg. Calculate the mass of  $1.5 \times 10^{-3} \text{ m}$  aqueous solution required for the above dose.

### Answer 1.29:

The molar mass of Nalorphene ( $\text{C}_{19}\text{H}_{21}\text{NO}_3$ ) is given as:

$$19 \times 12 + 21 \times 1 + 1 \times 14 + 3 \times 16 = 311 \text{ g mol}^{-1}$$

In  $1.5 \times 10^{-3} \text{ m}$  aqueous solution of nalorphene, 1 kg (1000 g) of water contains  $1.5 \times 10^{-3} \text{ mol}$

$$= 1.5 \times 10^{-3} \times 311 \text{ g} = 0.4665 \text{ g}$$

Therefore, total mass of the solution =  $(1000 + 0.4665) \text{ g} = 1000.4665 \text{ g}$

This implies that the mass of the solution containing 0.4665 g of nalorphene is 1000.4665 g.

Therefore, mass of the solution containing 1.5 mg of nalorphene is:  $\frac{1000.4665 \times 1.5 \times 10^{-3}}{0.4665} \text{ g} = 3.22 \text{ g}$

Hence, the mass of aqueous solution required is 3.22 g.

# Answer Of Chemistry Class 12

**Léo Pomar**



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web nov 16 2019 from the hebrew name אֵיתָן eitan meaning solid enduring firm in the old testament this name is borne by a few minor characters including the wise man ethan the ezrahite supposedly the author of psalm 89

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web jul 21 2021 it s a name with enduring strength and popularity appearing everywhere from the mission impossible movies to the well known furniture chain origin ethan comes from the hebrew name eitan and is featured in the old testament gender ethan is most often a boy name

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web nov 9 2023 the name ethan is a boy s name of hebrew origin meaning strong firm ethan is a name that succeeds in being at once classic and fashionable serious and cheery strong and sensitive ethan which was given a big boost via the name of the tom cruise character in the mission impossible film series is still popular but dipping a bit

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