

# Chemistry, Part 1

## Module 14 Notes



### Big Question

How are metals used in cement, fertilizers, pigments, and products for chemical reactions?

### Big Idea

Using dimensional analysis, you can convert from units of a substance to quantities of other substances. A metric conversion and dimensional analysis can be substituted by an international analysis with chemical equations.

### Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use stoichiometric equations for describing processes of reactions.
- Convert between moles, mass, and molecules.
- Describe uses of stoichiometry in industry.

### Mastering Key Skills

#### Learning Objective

How do you balance a chemical equation?

Fill in the blanks:

1.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  (Balanced)
2.  $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$  (Balanced)
3.  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  (Balanced)

#### Chemical Reactions and Equations

For the first activity, see your chemistry textbook.

Stoichiometric law is concerned with quantitative changes.

Mass conservation:  $\text{Mass}_R = \text{Mass}_P$  (Closed System)

Volume:  $V_R = V_P$  (Open System)

Molar conservation:  $n_R = n_P$  (Closed)

### Big Idea

#### Learning Objective

How do you use mass ratios in chemical equations?

Fill in the blanks:

1. 1 mole of oxygen is equal to weight \_\_\_\_\_ grams or 16 \_\_\_\_\_.
2. The mass of 1 mole of calcium is 40g. How many atoms are present in 1 mole of calcium? \_\_\_\_\_.

# Answers For Chemistry Module 14

**JS Bruner**



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