



Name:

Date:

Student Exploration: Calorimetry Lab

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- The Latin word *calor* means "heat," and *meter* comes from the Greek word meaning "to measure." What do you think a **calorimeter** does?

measuring the heat of chemical reactions or physical changes as well as heat capacity

- Where have you heard the word **calorie** before? What do you think a calorie is?

calorie, a unit of energy or heat variously defined. the calorie was originally defined as the amount of heat required at a pressure of 1 standard atmosphere to raise the temperature of 1 gram of water 1 degree celsius.

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

- What was the **Final temperature** of the copper and the water?
- How much did the temperature of the copper change?
- How much did the temperature of the water change?

34.96

down 55.04

down 165.04

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Answer Sheet For Calorimetry Lab

Ying-Ying Zheng



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Table of Contents Answer Sheet For Calorimetry Lab

1. Understanding the eBook Answer Sheet For Calorimetry Lab
 - The Rise of Digital Reading Answer Sheet For Calorimetry Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet For Calorimetry Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet For Calorimetry Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet For Calorimetry Lab
 - Personalized Recommendations
 - Answer Sheet For Calorimetry Lab User Reviews and Ratings
 - Answer Sheet For Calorimetry Lab and Bestseller Lists
5. Accessing Answer Sheet For Calorimetry Lab Free and Paid eBooks
 - Answer Sheet For Calorimetry Lab Public Domain eBooks
 - Answer Sheet For Calorimetry Lab eBook Subscription Services

- Answer Sheet For Calorimetry Lab Budget-Friendly Options
- 6. Navigating Answer Sheet For Calorimetry Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Sheet For Calorimetry Lab Compatibility with Devices
 - Answer Sheet For Calorimetry Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet For Calorimetry Lab
 - Highlighting and Note-Taking Answer Sheet For Calorimetry Lab
 - Interactive Elements Answer Sheet For Calorimetry Lab
- 8. Staying Engaged with Answer Sheet For Calorimetry Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet For Calorimetry Lab
- 9. Balancing eBooks and Physical Books Answer Sheet For Calorimetry Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet For Calorimetry Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Sheet For Calorimetry Lab
 - Setting Reading Goals Answer Sheet For Calorimetry Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Sheet For Calorimetry Lab
 - Fact-Checking eBook Content of Answer Sheet For Calorimetry Lab
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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