



# General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

## Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{P_1}{P_2} = \frac{r_1}{r_2}$$

$$P_{\text{solution}} = X_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = X_{\text{solute}} P_{\text{solvent}}^{\circ}$$

## Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \text{ (at } 25^{\circ}\text{C)}$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592V}{n} \log(K) \text{ (at } 25^{\circ}\text{C)}$$

USE ONE:

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} - E_{\text{anode(re.d.)}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} + E_{\text{anode(ox.)}}^{\circ}$$

## Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surroundings}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{surroundings}}$$

$$\Delta S_{\text{rxn}}^{\circ} = \Sigma S_{\text{products}}^{\circ} - \Sigma S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{rxn}}^{\circ} = \Sigma \Delta G_{\text{products}}^{\circ} - \Sigma \Delta G_{\text{reactants}}^{\circ}$$

## Concentrations

$$\text{molarity} = \frac{\text{moles}_{\text{solute}}}{\text{volume}_{\text{solution}}(\text{L})}$$

$$\% \text{ by mass} = \frac{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}}}{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}} + \text{mass}_{\text{solvent}}} \times 100\%$$

$$\text{parts per million} = \frac{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}}}{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}} + \text{mass}_{\text{solvent}}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}}}{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}} + \text{mass}_{\text{solvent}}} \times 10^9$$

## Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_a + pK_b = 14.00 \text{ (at } 25^{\circ}\text{C)}$$

$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$$pOH = pK_b + \log \frac{[BH^+]}{[B]}$$

$$\% \text{ ionization} = \frac{[H_3O^+]}{[HA]_{\text{initial}}} \times 100\%$$

$$pK_a = -\log(K_a)$$

$$pK_b = -\log(K_b)$$

$$K_a = 10^{-pK_a}$$

$$K_b = 10^{-pK_b}$$

## Equilibria

$$Q_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[A]^x[B]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k} \text{ (1st order reactions only)}$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$



# Answer Sheet General Chemistry Equations

**Marco Casella**



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## **Table of Contents Answer Sheet General Chemistry Equations**

1. Understanding the eBook Answer Sheet General Chemistry Equations
  - The Rise of Digital Reading Answer Sheet General Chemistry Equations
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet General Chemistry Equations
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Sheet General Chemistry Equations
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet General Chemistry Equations
  - Personalized Recommendations
  - Answer Sheet General Chemistry Equations User Reviews and Ratings
  - Answer Sheet General Chemistry Equations and Bestseller Lists
5. Accessing Answer Sheet General Chemistry Equations Free and Paid eBooks
  - Answer Sheet General Chemistry Equations Public Domain eBooks
  - Answer Sheet General Chemistry Equations eBook Subscription Services
  - Answer Sheet General Chemistry Equations Budget-Friendly Options
6. Navigating Answer Sheet General Chemistry Equations eBook Formats

- ePub, PDF, MOBI, and More
  - Answer Sheet General Chemistry Equations Compatibility with Devices
  - Answer Sheet General Chemistry Equations Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Answer Sheet General Chemistry Equations
    - Highlighting and Note-Taking Answer Sheet General Chemistry Equations
    - Interactive Elements Answer Sheet General Chemistry Equations
  8. Staying Engaged with Answer Sheet General Chemistry Equations
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Answer Sheet General Chemistry Equations
  9. Balancing eBooks and Physical Books Answer Sheet General Chemistry Equations
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Answer Sheet General Chemistry Equations
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Answer Sheet General Chemistry Equations
    - Setting Reading Goals Answer Sheet General Chemistry Equations
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Answer Sheet General Chemistry Equations
    - Fact-Checking eBook Content of Answer Sheet General Chemistry Equations
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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