

## Lab #2: Avogadro Goes to Court

The purpose is to use dimensional analysis in order to find the cost of a single aluminum atom in the roll of aluminum foil.

### Given information:

- Measurement of my piece of foil:  $30.48\text{ cm} \times 27.31\text{ cm} = 832.41\text{ cm}^2$
- Mass of my foil:  $2.79\text{ grams of aluminum}$
- Cost of Aluminum foil roll:  $\$2.99$
- Total area of the roll:  $6.56\text{ m}^2$
- Aluminum atomic mass:  $26.98\text{ g}$
- Avogadro's number:  $6.02 \times 10^{23}$

First, we need to get the area to mass ratio of the aluminum foil so that we already have all the information we need.

$$\text{Area to mass} = \frac{\text{area}}{\text{mass}} = \frac{(\text{length} \times \text{width})}{\text{mass}} = \frac{30.48\text{ cm} \times 27.31\text{ cm}}{2.79\text{ g}} = 294.3\text{ cm}^2/\text{g}$$

Convert to square meters:

$$\frac{294.3\text{ cm}^2}{\text{g}} \times \frac{1\text{ m}}{100\text{ cm}} \times \frac{1\text{ m}}{100\text{ cm}} = 0.02943\text{ m}^2/\text{g}$$

Use the total area of foil and determine how many grams a roll of foil has.

$$6.56\text{ m}^2 \times \frac{1\text{ g}}{0.02943\text{ m}^2} = 223.0\text{ g}$$

Convert 223.0 g into atoms.

$$223.0\text{ g}_{\text{Al}} \times \frac{1\text{ mol}_{\text{Al}}}{26.98\text{ g}_{\text{Al}}} \times 6.022 \times 10^{23}\text{ atoms}_{\text{Al}} = 4.93 \times 10^{24}\text{ atoms}_{\text{Al}}$$

Lastly, divide the price of the whole roll by the number of atoms to find the price per atom.

$$\frac{\$2.99}{4.93 \times 10^{24}}\text{ atoms}_{\text{Al}} = \$2.44 \times 10^{-25}\text{ per aluminum atom}$$

# Avogadro Goes To Court Lab Answer

**Jiyuan Zhang**



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