

CHAPTER 17 - PRINCIPLES OF CHEMICAL REACTIVITY - THE CHEMISTRY OF ACIDS AND BASES

I. ACIDS AND BASES: A REVIEW

1. Arrhenius:

- When acids dissolve, the $[H^+]$ of H^+ increases $[H^+]$
- Bases increase their $[OH^-]$ when dissolved in H_2O .

2. Brønsted-Lowry:

- A Brønsted acid is a proton (H^+) donor
- A Brønsted base is a proton (H^+) acceptor.

II. THE BRØNSTED-LOWRY CONCEPT OF ACIDS AND BASES EXTENDED.

1. There are VARIETIES OF BRØNSTED ACIDS + BASES



2. Monoprotic acid: Donate ONE proton.

3. Polyprotic acid: Donate TWO OR MORE protons.

a. H_2SO_4 and HSO_4^- :



4. Some molecules are acids/bases = amphiprotic

| ACID | Amphiprotic | BASE |
|-------------|----------------------------|---------------|
| H_2S | HS^- | S^{2-} |
| H_3PO_4 | $H_2PO_4^- / H_2PO_4^{2-}$ | PO_4^{3-} |
| H_2CO_3 | HCO_3^- | CO_3^{2-} |
| $H_2C_2O_4$ | $HC_2O_4^-$ | $C_2O_4^{2-}$ |

A. Conjugate Acid-Base Pair

- Every reaction bit on acid and base involves 2 CA-B PAIRS



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