

Задача 1
Уравновешивание химических реакций

1) Дано:
 $C_{HNO_3} = 0,01 \text{ м}$
 $K_a = 7,2 \cdot 10^{-4}$
 $pH = ?$

Решение:
 Для слабой кислоты:
 $[H^+] = \sqrt{K_a \cdot C_{HNO_3}}$
 $= \sqrt{7,2 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-3} \text{ моль/л}$
 $pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-3} = 3 - \lg 2,8 = 2,55$
 $= 2,55$

2) Дано:
 $IP_{Ca(OH)_2} = 6,3 \cdot 10^{-6}$
 $pH = ?$

Решение:
 $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$
 $IP = [Ca^{2+}] [OH^-]^2 = 6 \cdot (2S)^2 = 12S^3$

S - предельная концентрация
 $S = \sqrt[3]{\frac{IP}{12}} = \sqrt[3]{\frac{6,3 \cdot 10^{-6}}{12}} = 1,15 \cdot 10^{-2} \text{ моль/л}$
 $[OH^-] = \frac{S}{2} = \frac{1,15 \cdot 10^{-2}}{2} = 5,75 \cdot 10^{-3} \text{ моль/л}$
 $[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{5,75 \cdot 10^{-3}} = 1,75 \cdot 10^{-12}$
 $pH = -\lg [H^+] = -\lg 1,75 \cdot 10^{-12} = 12 - \lg 1,75 = 11,25$
 $= 11,25$

3) Дано:
 $V_{CaCl_2} = 10 \text{ мл}$
 $C_{CaCl_2} = 0,1 \text{ М}$
 $V_{Na_2CO_3} = 10 \text{ мл}$
 $C_{Na_2CO_3} = 0,1 \text{ М}$
 $IP_{CaCO_3} = 10^{-9}$

Решение:
 после смешивания:
 $C_{CaCl_2} = 0,1 \cdot \frac{10}{20} = 0,05 \text{ моль/л}$
 $C_{Na_2CO_3} = 0,1 \cdot \frac{10}{20} = 0,05 \text{ моль/л}$
 или предельная концентрация:
 $C_{Ca^{2+}} = C_{CaCl_2} = 0,05 \text{ моль/л}$
 $C_{CO_3^{2-}} = C_{Na_2CO_3} = 0,05 \text{ моль/л}$
 $[Ca^{2+}] [CO_3^{2-}] = 0,05 \cdot 0,05 = 2,5 \cdot 10^{-3} > 10^{-9}$

3) $PbCl_2 \rightleftharpoons Pb^{2+} + 2Cl^-$
 $IP_{PbCl_2} = 1,7 \cdot 10^{-5}$

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