

Задача 1
Уравнение и моляр. констант

① Дано:
 $C_{\text{HAc}} = 0,01 \text{ м}$
 $K_a = 7,2 \cdot 10^{-4}$
 $\text{pH} = ?$

Решение:
 Для слабой кислоты:
 $[H^+] = \sqrt{K_a \cdot C_{\text{HAc}}}$
 $= \sqrt{7,2 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-3} \text{ моль/л}$
 $\text{pH} = -\lg [H^+] = -\lg 2,8 \cdot 10^{-3} = 3 - \lg 2,8 = 2,55$
 $= 2,55$

② Дано:
 $\text{p}K_{\text{Ca}^{2+}} = 6,3 \cdot 10^6$
 $\text{pH} = ?$

Решение:
 $\text{Ca}^{2+} + 2\text{OH}^- \rightleftharpoons \text{Ca}(\text{OH})_2$
 $K_{\text{пр}} = [\text{Ca}^{2+}] [\text{OH}^-]^2 = 9 \cdot 10^{-7}$

S - растворимость $\text{Ca}(\text{OH})_2$
 $S = \sqrt[3]{\frac{K_{\text{пр}}}{4}} = \sqrt[3]{\frac{9 \cdot 10^{-7}}{4}} = 1,15 \cdot 10^{-2} \text{ моль/л}$
 $[\text{OH}^-] = \frac{S}{2} = \frac{1,15 \cdot 10^{-2}}{2} = 5,75 \cdot 10^{-3} \text{ моль/л}$
 $[\text{H}^+] = \frac{10^{-14}}{[\text{OH}^-]} = \frac{10^{-14}}{5,75 \cdot 10^{-3}} = 1,75 \cdot 10^{-12}$
 $\text{pH} = -\lg [H^+] = -\lg 1,75 \cdot 10^{-12} = 12 - \lg 1,75 = 11,25$
 $= 11,25$

③ Дано:
 $V_{\text{HCl}} = 10 \text{ мл}$
 $C_{\text{HCl}} = 0,1 \text{ М}$
 $V_{\text{NaOH}} = 10 \text{ мл}$
 $C_{\text{NaOH}} = 0,1 \text{ М}$
 $\text{p}K_{\text{PbCl}_2} = 1,6 \cdot 10^4$

Решение:
 после смешивания:
 $C_{\text{HCl}} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$
 $C_{\text{NaOH}} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$
 или гидролизированной кислоты:
 $C_{\text{H}^+} = C_{\text{OH}^-} = 0,05 \text{ моль/л}$
 $C_{\text{Pb}^{2+}} = C_{\text{Cl}^-} = 0,05 \text{ моль/л}$
 $[Pb^{2+}] [Cl^-]^2 = 6,05 \cdot 10^{-5} \cdot 10^{-4} = 6,05 \cdot 10^{-9}$

④
 $\text{PbCl}_2 \rightleftharpoons \text{Pb}^{2+} + 2\text{Cl}^-$
 $K_{\text{пр}} = 6,05 \cdot 10^{-9}$

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Ashy Kumren, Chwee Neo Lee



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