

Trial 1:

Mass of test tube and clamp:	Mass of test tube, clamp and baking soda	Mass of baking soda
34.52g	35.74g	1.22g NaHCO_3

Let's review steps:

Convert the mass of baking soda to moles

$$1.22 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84 \text{ g NaHCO}_3} = 0.0145 \text{ mol NaHCO}_3$$

Using the balanced equation, how many moles of carbon dioxide will be produced?



Convert the moles of carbon dioxide to grams of carbon dioxide. (theoretical mass of carbon dioxide.)

Use the ideal gas law to determine how many liters this amount of CO_2 will take up at STP.

Analysis Of Baking Soda Lab Answers

JL Elias



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