

1. **Strategically** balance the hydroxide and sulfide ions (use moles) of a solution and identify each component. This solution is given the percentage composition but also the percent amount of each in a substance that is dissolved in the solution (mass % solution). There is the solvent and solute in the solution as well as the solute.

2. What is the composition of a mixture in each of the following aqueous solutions?

a) 4.20 M Na_2SO_4 $\text{mass}(\%) = \frac{4.20 \text{ mol Na}_2\text{SO}_4}{4.20 \text{ mol Na}_2\text{SO}_4 + 1 \text{ mol H}_2\text{O}} \times 100\% = 82.0\% \text{ Na}_2\text{SO}_4$

b) 0.20 M Na_2SO_4 $\text{mass}(\%) = \frac{0.20 \text{ mol Na}_2\text{SO}_4}{0.20 \text{ mol Na}_2\text{SO}_4 + 1 \text{ mol H}_2\text{O}} \times 100\% = 17.2\% \text{ Na}_2\text{SO}_4$

c) 0.20 M Na_2SO_4 $\text{mass}(\%) = \frac{0.20 \text{ mol Na}_2\text{SO}_4}{0.20 \text{ mol Na}_2\text{SO}_4 + 1 \text{ mol H}_2\text{O}} \times 100\% = 17.2\% \text{ Na}_2\text{SO}_4$

3. A solution is prepared by dissolving 20.0 g of sodium sulfate in enough water to prepare 250. mL of solution.

- a) What is the molarity of sodium sulfate in the solution?

$$\text{M}_\text{Na}_2\text{SO}_4 \text{ in } 250. \text{ mL } = 2 \text{ mol Na}_2\text{SO}_4 = 20.0 \text{ g Na}_2\text{SO}_4 \times \frac{1 \text{ mol Na}_2\text{SO}_4}{142.0 \text{ g Na}_2\text{SO}_4} = 0.141 \text{ mol Na}_2\text{SO}_4 = 0.141 \text{ mol Na}_2\text{SO}_4 / 0.250 \text{ L} = 0.563 \text{ mol/L}$$

$$\text{M}_\text{Na}^+ \text{ in } 250. \text{ mL } = 2 \text{ mol Na}^+ = 20.0 \text{ g Na}_2\text{SO}_4 \times \frac{2 \text{ mol Na}^+}{142.0 \text{ g Na}_2\text{SO}_4} = 0.282 \text{ mol Na}^+ = 0.282 \text{ mol Na}^+ / 0.250 \text{ L} = 1.13 \text{ mol/L}$$

$$\text{M}_\text{SO}_4^{2-} \text{ in } 250. \text{ mL } = 1 \text{ mol SO}_4^{2-} = 20.0 \text{ g Na}_2\text{SO}_4 \times \frac{1 \text{ mol SO}_4^{2-}}{142.0 \text{ g Na}_2\text{SO}_4} = 0.141 \text{ mol SO}_4^{2-} = 0.141 \text{ mol SO}_4^{2-} / 0.250 \text{ L} = 0.563 \text{ mol/L}$$

- b) What are the molarities of the sodium and sulfate ions in the solution?

$$\text{Na}^+ \text{ in } 250. \text{ mL } = 2 \text{ mol Na}^+ = 20.0 \text{ g Na}_2\text{SO}_4 \times \frac{2 \text{ mol Na}^+}{142.0 \text{ g Na}_2\text{SO}_4} = 0.282 \text{ mol Na}^+ = 0.282 \text{ mol Na}^+ / 0.250 \text{ L} = 1.13 \text{ mol/L}$$

$$\text{SO}_4^{2-} \text{ in } 250. \text{ mL } = 1 \text{ mol SO}_4^{2-} = 20.0 \text{ g Na}_2\text{SO}_4 \times \frac{1 \text{ mol SO}_4^{2-}}{142.0 \text{ g Na}_2\text{SO}_4} = 0.141 \text{ mol SO}_4^{2-} = 0.141 \text{ mol SO}_4^{2-} / 0.250 \text{ L} = 0.563 \text{ mol/L}$$

- c) How many moles of sodium ions are present in 17 mL of the solution?

$$\text{mol Na}^+ = \frac{17 \text{ mL}}{1000 \text{ mL}} \times 1.13 \text{ mol Na}^+ = 0.0192 \text{ mol Na}^+$$

4. A solution of Na_2SO_4 which has a volume of 750 mL, contains 200.0 moles of potassium ions. What is the molarity of the Na_2SO_4 solution?

$$\frac{200.0 \text{ mol K}^+}{2 \text{ mol K}^+} = \frac{200.0 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 100.0 \text{ mol Na}_2\text{SO}_4$$

5. What are the two major hydrolytic products?

Hydrolysis reaction of weak acids in the solution and giving the weak products in the solution.

6. Explain the hydrolytic effect.

This OH⁻ groups of negative compounds are stronger and release very easily with water. Compounds with more OH⁻ groups release or easily with water that they are called hydrolytic. When a hydrolytic compound is placed in water, the water molecules react to get a stronger than compound with the compound. As a result hydrolytic compounds are not stable in water.

7. Rank the following acids in order of their ability to H_2O or $\text{Cl}_2\text{H}_2\text{O}$.

a) HCl (aq) b) H_2SO_4 (aq) c) HCl_2O_2 (aq) d) H_2O (aq)

Hydrogen bonding substance (H⁺ and OH⁻) and weak substances (HCl₂O₂) are more soluble in water than it is stronger substance, while weaker substances (H₂O) are more soluble in weaker substance.

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