

- 2013**
- Q.1 Define Acid, Base and Salt. Also give one example of each.
 Q.2 Calculate the pH and pOH of a 0.1 M Nitric Acid solution.
- 2014**
- Q.1 Define double salt and write the formulae of:
 (i) Potash Alum (ii) Carnallite
- Q.2 Define:
 (i) Standard solution (ii) Acidity of a base (iii) Basicity of acid
- Q.3 Define water of crystallization, write names and chemical formulae of two hydrated salts.
 Q.4 Describe with balanced chemical equations Solvay process for Industrial preparation of Anhydrous Sodium Carbonate.
- 2015**
- Q.1 State Bronsted-Lowry theory of acid and basis with one example of each.
 Q.2 Write chemical formula of the following hydrated salt with their water of crystallization.
 (i) Copper sulphate (ii) Sodium carbonate (iii) Magnesium sulphate
- Q.3 Define neutralization? Write down chemical equation of the reaction of NaOH with HCl and CH_3COOH .
 Q.4 Define pH.
- 2016**
- Q.1 State the Lewis concept about acids and bases and explain it with an example.
 Q.2 Describe the extraction of Sulphur by Frasch process.
- 2017**
- Q.1 Describe the Lewis concept about acids and bases.
 Q.2 Define salts and describe three types of salts with examples.

ACID, BASES AND SALTS

- Q.1 (a) Define the following terms
- | | | |
|--------------------------|-----------------------|--------------------|
| i. Acids | ii. Bases | iii. Strong acids |
| iv. Weak Acids | v. Strong bases | vi. Weak bases |
| vii. Basicity of an acid | viii. Acidity of base | ix. Neutralization |
| x. Titration | xi. pH | xii. pOH |
| xiii. Standard Solution | | |
- Q.1 (b) Write the chemical formulae of:
 (i) Caustic soda (ii) Soda ash (iii) Baking soda

Ans.1 a. **Acids:** Acids can be defined as
 "The compounds which produce hydrogen ions (H^+) in aqueous solution are called acids".
 The word acid is derived from a Latin word *acidus*, which means sour. For example HCl, HBr, HI

Bases: Bases can be defined as
 "The compounds which produce hydroxide ions (OH^-) in aqueous solution are called bases". For example NaOH, KOH

Strong Acids:
 Acids which ionize completely in an aqueous solution are strong acids. For example H_2SO_4 , HCl, HNO_3

Weak Acids:
 Acids which do not ionize completely in an aqueous solution are weak acids. For example CH_3COOH , H_2CO_3

Strong Bases:
 Bases which ionize completely in aqueous solution are called as strong bases. For example NaOH, KOH

Ch Acids Bases Salts Study Guide Answer

Katrin Zwirgmaier



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