

**CHEM 1411 – General Chemistry I**  
**Practice Problems, Chapters 1–3**

**Chapter 1 – Chemistry: The Study of Change**

- 1.** Element, compound, homogeneous mixture (solution), or heterogeneous mixture:
- |                                       |                       |   |
|---------------------------------------|-----------------------|---|
| a) orange juice<br>(freshly-squeezed) | b) brass              | c) 0.9% saline (NaCl) solution                |
| d) garden soil                        | e) room air           | f) methane gas                                |
| g) sodium metal                       | h) N <sub>2</sub> gas | i) Cu(NO <sub>3</sub> ) <sub>2</sub> crystals |
- 2.** Define (some of these terms are found in Chapters 2 and 3):
- |                         |                                |
|-------------------------|--------------------------------|
| a) element              | b) compound                    |
| c) solution             | d) physical change             |
| e) chemical change      | f) law of conservation of mass |
| g) intensive properties | h) extensive properties        |
| i) precision            | j) accuracy                    |
| k) isotopes             | l) stoichiometry               |
| m) the mole             | n) Avogadro's number           |
| o) limiting reactant    |                                |
- 3.** Convert
- 5 feet 10 inches to meters
  - 55 miles per hour to cm per second
  - 5.0 quarts to liters
  - 1.00 ft<sup>2</sup> to cm<sup>2</sup>
  - 32.06 amu to g
  - 25 g of carbon to moles of carbon
  - 1.00 g of gold to number of atoms of gold
  - 25 °F to Kelvins
- 4.** How many significant figures are in the following numbers?
- |             |  |               |
|-------------|--|---------------|
| a) 3.1416   | b) 6.022 × 10 <sup>23</sup> atoms/mole | c) 0.000104 g |
| d) 6.0 g/mL | e) 1.00794 amu                         | f) 5000 mi    |

# Chapter 1 Chemistry Test

**J Dewey**



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