

## Chemistry Student Edition - Basic Answer Key

### Chapter 19: Chemical Equilibrium

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#### The Nature of Chemical Equilibrium

##### Questions

- At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction?
- What does the law of mass action say about any reaction at equilibrium?
- Why are solids and liquids left out of the equilibrium constant expression?
- Explain why the equilibrium constant expression as derived from concentrations can be converted to the equilibrium constant expression as derived from partial pressures.
- Construct balanced chemical equations and write the equilibrium constant expression (either  $K_p$  or  $K_c$ ) for each of the following processes:
  - Sodium metal reacts vigorously with oxygen gas to produce sodium oxide.
  - Aluminum oxide decomposes to form aluminum metal and oxygen gas.
  - Nitrogen gas and oxygen gas react to form nitrogen monoxide gas.
- Give an example of a chemical equilibrium that exists in our daily lives.
- Write a  $K_{eq}$  expression for each of the following processes:
  - Solid calcium carbonate is heated and decomposes into solid calcium oxide and carbon dioxide gas.
  - Hydrogen gas ( $H_2$ ) and solid iodine ( $I_2$ ) are combined to form hydrogen iodide gas ( $HI$ ).

##### Answers

- At equilibrium, the rate of the forward reaction equals the rate of the reverse reaction.
- The ratio of product concentrations to reactant concentrations (along with the appropriate exponents) is equal to a constant.
- Solids and liquids have essentially constant densities, which means that writing them is not very useful.
- Partial pressures are related to concentrations by a factor of  $RT$ .
- 



$$K_c = \frac{[Na_2O_{(s)}]^2}{[Na_{(s)}]^2 [O_{2(g)}]}$$

- a) heterogeneous equilibria



$$K_c = \frac{[Al_{(s)}]^4 [O_{2(g)}]^3}{[Al_2O_{3(s)}]^2}$$

- b) heterogeneous equilibria

# Answer Key Ch 19 Chemistry

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