



## SPECIFIC HEAT WORKSHEETS

$Q = mc\Delta T$ ,  
where  $Q$  = heat energy,  $m$  = mass, and  
 $\Delta T$  = change in temp.  
Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ .



Show all work and proper units.

- A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.  
 $c = q / (m \cdot \Delta T)$   
 $c = 1086.75 \text{ J} / (15.75 \text{ g} \times 150 \text{ }^\circ\text{C})$   
 $= 0.46 \text{ J/g }^\circ\text{C}$
- How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?  
 $q = mc\Delta T$   
 $q = (10.0\text{g})(0.90\text{J/g}^\circ\text{C})(33^\circ\text{C})$   
 $q = 297 \text{ Joules}$
- Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.  
 $c = q / (m \cdot \Delta T)$   
 $c = 67,500 \text{ J} / (1500.0 \text{ g} \times 25 \text{ }^\circ\text{C})$   
 $= 1.8 \text{ J/g }^\circ\text{C}$
- 100.0 g of 4.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.  
 $q = mc\Delta T$   
 $q = (100 \times 4.18 \times 33) \text{ J}$   
 $q = 13,794 \text{ J}$
- Calculate the heat required to raise 0.6 Kg of sand from 30°C to 90°C? (Specific Heat of sand = 830 J/Kg°C)  
 $q = mc\Delta T$   
 $q = 830 \text{ J/Kg}^\circ\text{C} \times 0.6 \text{ Kg} \times 60^\circ\text{C}$   
 $q = 29880 \text{ J.}$

# Answer Key Heat

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