

Question	Answer	Marks	Guidance
1 (a) (i)	94 protons	1	This is the proton number Z .
1 (a) (ii)	145 neutrons	1	Number of neutrons is found from (nucleon number - proton number) = $(A - Z) = 239 - 94$
1 (a) (iii)	93 electrons	1	A neutral atom would have 94 electrons (= number of protons Z). This ion has a charge of $+e$ ($+1.6 \times 10^{-19}$ C), showing that there is one electron missing.
1 (b)	isotopes have same number of protons (or the same Z) but a different number of neutrons (or nucleons, or a different A)	1 1	You need to remember how isotopes are defined.
2 (a) (i)	$X =$ a neutrino ${}_{29}^{64}\text{Cu} \rightarrow {}_{28}^{64}\text{Ni}$ ${}_{-1}^0\beta + \nu$	1 1 1	The proton number and nucleon number must balance on both sides of the equation.
2 (a) (ii)	The copper nucleus has the greater specific charge because it has a charge of $+29e$ whereas the nickel nucleus has a charge of $+28e$ and they both have the same mass (to 2 significant figures).	1 1	The proton number gives you the total number of positively-charged protons in the nucleus.
2 (a) (iii)	specific charge = $\frac{29e}{64m_u}$ $= \frac{29 \times 1.6 \times 10^{-19}}{64 \times 1.67 \times 10^{-27}}$ $= 4.3 \times 10^7 \text{ C kg}^{-1}$	1 1 1	To find the specific charge, simply divide the charge by the mass of the nucleus. Each proton carries a positive charge of $+e$.
2 (b) (i)	udd	1	It's useful to know this.
2 (b) (ii)	An up quark in a proton changes into a down quark so the proton becomes a neutron. At the same time, the up quark emits a W^+ boson when it changes into a down quark and the W^+ boson decays into a β^+ particle and a neutrino.	1 1	Make sure you don't forget the neutrino in this kind of interaction.
3 (a)	$\pi^+ + p \rightarrow n + \pi^+ + K^+$	1	All that is required is to translate the words of the question into symbols.

Atomic Nuclei Section 1 Answers

Michael Rajnik



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