

ATOMIC STRUCTURE

RUTHERFORD'S MODEL

- Discovered that an atom is mostly made up of empty space
- Center of the atom is a small dense positively charged nucleus
- This attraction holds the negatively charged particles in the atom

BOHR'S MODEL

- There are 2 main pieces of an atom:
- Nucleus
- Electrons

THE EARLY MODEL OF THE ATOM



WAVE-MECHANICAL MODEL

- Electrons are moving so fast that they look like a cloud. Orbitals are regions of most likely electron location



WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
 - PROTONS - positively charged particles
 - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
 - Electrons DO move around the nucleus in different patterns



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1 amu
neutron	center (in the nucleus)	neutral (no charge)	1 amu
electron	around the nucleus in clouds or orbitals	negative	Almost equal to 1/1836

IONS

- If an atom gains or loses an electron, it is called an ion
- This gives the atom a charge:
 - Lose an electron, the atom becomes positive
 - Gain an electron, the atom becomes negative

ISOTOPES

- An isotope is when an atom is proton or neutron
- An isotope is still the same element, it is just a little different from every other atom of the same element
 - slightly different mass

Atomic Structure Anticipation Guide

M Planty



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Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Chemistry, Study Guide James E. Brady,John R. Holum,1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

Essential Concepts of

Chemistry Study Guide James R. Braun, Sherman, 1999 *Library of Congress Catalog: Motion Pictures and Filmstrips*
Library of Congress, 1968 Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems **Study Guide for General Chemistry for Colleges** Herman Thompson Briscoe, 1936 **Study Guide for E. Russell Hardwick's Introduction to Chemistry** Joan Bouillon, 1984 Student Study Guide to Accompany Human Biology Sylvia S. Mader, 2003-08 This best selling text emphasizes the relationship between humans and other living things Intended for an introductory course this text provides students with a firm grasp of how their bodies function and how the human population can become more fully integrated into the biosphere An Online Learning Center tied directly to the text via icons will direct students to activities or animations that gives a visual example of difficult processes as well as Working Together boxes to emphasize homeostasis **Study Guide and Student Solutions Manual to Accompany Physics for Scientists and Engineers, by Serway** Raymond A. Serway, Steven Van Wyk, 1996 *Student Study Guide to accompany Chemistry* Martin Silberberg, Libby Weberg, 2005-01-06 *Study Guide for Chemical Principles [by] Steven S. Zumdahl* Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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