

Thermochemistry Worksheet Answers

Useful Values:

Heat of Fusion for Water: $\Delta H_{\text{fus}} = 6 \text{ kJ/mol}$ **Common error: not converting to moles**
 Heat of Vaporization for Water: $\Delta H_{\text{vap}} = 45 \text{ kJ/mol}$ **Common error: same**
 1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K	
Water (ice)	2.05	0.490	Common error here: using moles instead of grams.
Water (liquid)	4.18	1.00	
Water (gas)	2.08	0.498	

- How much energy does it take to warm the following substances the listed amount?
 - 20g of ice from -45°C to -15°C ? **294 calories / 1230 Joules**
 - 20g of water from 25°C to 55°C ? **600 calories / 2508 Joules**
 - 20g of water vapor from 130°C to 160°C ? **299 calories / 1249 Joules**
- To melt ice, one must bring ice to the freezing point (0°C), and then thaw it into water. How much heat would it take to melt...
 - 30g of ice at 0°C ? **10 kJ (10,000 J)**
 - 30g of ice at -40°C ? **12.46 kJ (12,460 J)**
- To boil water, one must raise it to the boiling point (100°C), and then boil it into steam. How much heat would it take to boil...
 - 40g of water at 100°C ? **100 kJ (100,000 Joules)**
 - 40g of water at 90°C ? **101.7 kJ (101,672 Joules)**
 - 40g of water at 0°C ? **116.7 kJ (116,720 Joules)**
 - 40g of ice at 0°C ? **130.1 kJ (130,053 Joules)**
 - 40g of ice at -80°C ? **136.6 kJ (136,613 Joules)**
- If 700 calories of heat raise 400g of Phonium by 20°C , what is the specific heat of Phonium? **0.0875 cal/g/K**
- How much heat energy is stored in 2g of ice at -30°C ? (remember: use Kelvin) **238 cal / 996 J**
- How much heat energy is stored in 2g of Phonium at 70°C ? (Assume Phonium is a solid at 70°C) **1672 cal**
- If 100g of water at 10°C comes in contact with _____, what will the final temperature be? (Remember, you have the specific heat of Phonium from question 4.)
 - 100g of water at 90°C **50° C**
 - 100g of Phonium at 90°C (Hint: the water and Phonium end at the same temperature) **16.44° C**
Common error: 6.44 degrees is the temperature change, and the water gets warmer.
 - 20g of Phonium at 90°C **11.38° C**
 - 1g of Phonium at 90°C **10.07° C**
- If 450g of Phakelite at 100°C is added to 1000g of water at 20°C , the temperature of the water increases by 5° . What is the specific heat of Phakelite? **0.148 cal/g/K**
- If 450g of Nonsensium at 500°C is added to 1000g of water at 20°C , the temperature of the water increases by 25° . What is the specific heat of Nonsensium? **0.122 cal/g/K**

Answers Thermochemistry Problems Worksheet Number Two

Daniel C. Tofan



Answers Thermochemistry Problems Worksheet Number Two:

General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *Singlet Photodissociation of Ketene* Melonie Anne Teichert, 1999

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