Chapter 1

Introduction to Chemistry

Practice Problems

1.16 Complete the following statement: Data that have a small random error but otherwise fall in a narrow range are (a) accurate, (b) precise, or (c) neither.

1.26 What is the difference between a qualitative and a quantitative measurement?

1.54 Carry out each of the following conversions: (a) 25.5 m to km, (b) 36.3 km to m, (c) 487 kg to g, (d) 1.32 L to mL, (e) 55.9 dL to L, (f) 6251 L to cm3

1.60 Mercury has a density of 13.6 g/mL. What is the mass of 4.72 L of mercury?

1.84 A student was given two metal cubes that looked similar. One was 1.05 cm on an edge and had a mass of 14.32 grams; the other was 2.66 cm on a side and had a mass of 215.3 grams. How can the student determine if these two cubes of metal are the same material using only the data given?

1.90 A calibrated flask was filled to the 25.00-mL mark with ethyl alcohol and was found to have a mass of 19.7325 g. In a second experiment, 25.0920 g of metal beads were put into the container and the flask was again filled to the 25.00-mL mark. The total mass of the metal plus the alcohol was 43.0725 g. Describe how to determine the density of the metal sample.

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Peipei Pang

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