



CHEMISTRY CHEAT SHEET



EQUILIBRIUM

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Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c .

| Chemical Equilibrium | | | | | | | | | | | | | | |
|---|---|---|----------|----------------------------|----------------------------|-------------------------|---|-------------------------------|----------------------------|---|------------------|-------------|--|---------|
| <p>What is Equilibrium?</p> <p>A reversible reaction is one that can proceed in either direction. At equilibrium, the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium Constant K_c</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K_c is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products at equilibrium.</p> | <p>Equilibrium Constant K_c</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K_c is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products at equilibrium.</p> | <p>Graphical Representation of Equilibrium</p> | | | | | | | | | | | | |
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Jiyuan Zhang



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