

**EQUILIBRIUM**

- 1. The equilibrium constant, which value gives a weak description of which reaction is favored?
- 2. If a reaction is not at equilibrium and the product is favored.
- 3. Equilibrium is favored when the concentration of the reactants is higher than the concentration of the products.
- 4. **EQUILIBRIUM CONSTANT**
- 5. **EQUILIBRIUM CONSTANT**

**EMPIRICAL REPRESENTATIONS**



1

For the equilibrium  $A \rightleftharpoons B$ , the reaction of the rate of the forward (1) & reverse (2) reaction will have a given by



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**EMPIRICAL REPRESENTATIONS**

- 1. Equilibrium constant is constant for a reaction.
- 2. The reaction is favored.
- 3. **EQUILIBRIUM CONSTANT**
- 4. The reaction is favored when the concentration of the reactants is higher than the concentration of the products.
- 5. **EQUILIBRIUM CONSTANT**
- 6. The reaction is favored when the concentration of the reactants is higher than the concentration of the products.
- 7. **EQUILIBRIUM CONSTANT**
- 8. The reaction is favored when the concentration of the reactants is higher than the concentration of the products.
- 9. **EQUILIBRIUM CONSTANT**
- 10. The reaction is favored when the concentration of the reactants is higher than the concentration of the products.

2

Which of the given statements does not describe the equilibrium state properly?  
 (A) The equilibrium can be approached from either direction.  
 (B) The equilibrium can be imposed only if the system is an isolated system.  
 (C) The free energy change of constant pressure and temperature is zero.  
 (D) It is dynamic in nature.

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**CHEMICAL EQUILIBRIUM**

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3

In the given reaction,  $A + B \rightleftharpoons C$ , 2 moles each of A & B are present in 10 l of volume contained in flask. It was at 10°C. Calculate K for the reaction.

$K = 1/10$     $K = 1/40$     $K = 1/10$     $K = 1/20$

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**CHEMICAL EQUILIBRIUM**

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4

In a given temperature, the equilibrium constant for the reaction,  
 $NO_2 \rightleftharpoons \frac{1}{2} N_2O_4$  is  $K_1$  and for the reaction,  
 $2NO_2 \rightleftharpoons N_2O_4$  is  $K_2$ , respectively. If  $K_1$  is  $4 \times 10^{-4}$ , then  $K_2$  will be  
 (A)  $4 \times 10^{-4}$    (B)  $16 \times 10^{-4}$    (C)  $25 \times 10^{-4}$    (D)  $25 \times 10^{-2}$

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**HOMOGENEOUS EQUILIBRIUM**

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5

For the reaction,  
 $Fe_3O_4 + 3H_2 \rightleftharpoons 3H_2O + 3Fe$   
 (a)  $K_1 = K_2(RT)^3$    (b)  $K_1 = K_2(RT)^{-3}$   
 (c)  $K_1 = K_2(RT)^{-3}$    (d)  $K_1 = K_2(RT)^3$

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**LE EQUILIBRIUM CONSTANT**

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6

Which one of the following conditions will favour maximum formation of the product in the reaction  $2SO_2 + O_2 \rightleftharpoons 2SO_3$ ,  $\Delta H = -198 \text{ kJ}$   
 (A) Low temperature and high pressure   (B) High temperature and high pressure  
 (C) Low temperature and low pressure   (D) High temperature and low pressure

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**D Sidentop**



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## **Table of Contents Chemistry Chemical Equilibrium Study Guide Key**

1. Understanding the eBook Chemistry Chemical Equilibrium Study Guide Key
  - The Rise of Digital Reading Chemistry Chemical Equilibrium Study Guide Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chemical Equilibrium Study Guide Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chemical Equilibrium Study Guide Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chemical Equilibrium Study Guide Key
  - Personalized Recommendations
  - Chemistry Chemical Equilibrium Study Guide Key User Reviews and Ratings
  - Chemistry Chemical Equilibrium Study Guide Key and Bestseller Lists
5. Accessing Chemistry Chemical Equilibrium Study Guide Key Free and Paid eBooks
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  - Chemistry Chemical Equilibrium Study Guide Key Enhanced eBook Features
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    - Interactive Elements Chemistry Chemical Equilibrium Study Guide Key
  8. Staying Engaged with Chemistry Chemical Equilibrium Study Guide Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Chemical Equilibrium Study Guide Key
  9. Balancing eBooks and Physical Books Chemistry Chemical Equilibrium Study Guide Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Chemical Equilibrium Study Guide Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Chemical Equilibrium Study Guide Key
    - Setting Reading Goals Chemistry Chemical Equilibrium Study Guide Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Chemical Equilibrium Study Guide Key
    - Fact-Checking eBook Content of Chemistry Chemical Equilibrium Study Guide Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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