

## Chemistry Cumulative Exam Review 2024 Questions and Answers

A student writes the following statements about the relationship between mass defect and nuclear binding energy.

1. When an isotope's protons and neutrons come together to form a nuclide, a small amount of mass is converted to energy.
2. An equivalent amount of energy must be added to the nuclide to split apart its protons and neutrons.
3. Statement 1 describes the isotope's nuclear binding energy, and statement 2 describes the isotope's mass defect.

Which statement is correct?

(Atoms and the Atomic Theory) - C) 1 and 2 only

In which isotope is the strong nuclear force not strong enough to balance electrostatic forces between protons?

(Atoms and the Atomic Theory) - D) uranium-238

Which is a correct comparison between the modern quantum model and John Dalton's model of the atom? - C) The modern model recognizes particles in the atom, whereas Dalton's model does not.

What did J. J. Thomson observe when he applied electric voltage to a cathode ray tube in his famous experiment?

(Atoms and the Atomic Theory) - B) The glass opposite of the negative electrode started to glow.

The mass number of a chromium atom is 52 and it has 24 protons. How many neutrons does this atom have?

(Atoms and the Atomic Theory) - B) 28

Which of these properties is the best one to use for identification of an element?

(Atoms and the Atomic Theory) - B) the number of protons in the atomic nucleus

# Chemistry Cumulative Review For Chapter 4 5 Answer Key

**Julia Schneider**



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