

Gen Chem II Study Slam 2

Question 1: Answer the following:

1. How long will it take for the amount of carbon-10 in a 2.00g sample of pure carbon-10 to decrease to 2.00×10^{-4} g? The half-life of carbon-10 is 19.3 seconds. (First-order)

$$k = \frac{\ln(2)}{t_{1/2}} = \frac{\ln(2)}{19.3} = 0.03541736162$$

$$\ln\left(\frac{2.00 \times 10^{-4} \text{ g}}{2.00 \text{ g}}\right) = -kt \quad t = \frac{\ln\left(\frac{2.00 \times 10^{-4}}{2.00}\right)}{-0.03541736162} = 256.452342476 \text{ s} =$$

$$\boxed{256 \text{ s}}$$

2. A reaction ($C \rightarrow \text{products}$) is found to be first order. How long will it take (in minutes) to consume 85.0% of reactant C, if the rate constant is $6.94 \times 10^{-3} \text{ s}^{-1}$?

100% - 85.0% = 15.0% remaining

$$\ln(0.15) = -kt$$

$$t = \frac{\ln(0.15)}{-6.94 \times 10^{-3} \text{ s}^{-1}} = 273.36012837 \text{ s} = \frac{273.36012837}{60} = 4.5560021406 \text{ min}$$

$$= \boxed{4.6 \text{ min}}$$

3. A reaction ($D \rightarrow \text{products}$) is found to be second order with $k = 7.35 \times 10^{-4} \text{ M}^{-1} \text{ s}^{-1}$. If after 10.0 hours of reacting, the concentration of the reactant D is 0.0253M, what must have been its initial concentration?

$$\frac{1}{[A]} = \frac{1}{[A]_0} + kt \quad \frac{1}{[A]_t} = \frac{1}{[A]} - kt \quad \frac{1}{[A]_0} = \frac{1}{[A]} - 10 \times 3600 (7.35 \times 10^{-4})$$

$$\frac{1}{[A]_0} = \frac{1}{0.0253} - 26.46$$

$$\frac{1}{[A]_0} = 12.7554362745 = \frac{0.07839640913}{\text{M}}$$

$$= \boxed{0.0784 \text{ M}}$$

Chemistry Episode 2 answer Key

**Hannah Cassidy, Cheryl Pelteret, Julie
Penn**

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