



CHEMISTRY

CHEAT SHEET



NAMING SIMPLE COMPOUNDS

NAMING SIMPLE COMPOUNDS CHEAT SHEET

Naming Compounds with Two Elements
 (Name the more electronegative element first)

- EXAMPLE**
- CO2
 Carbon Dioxide (Naming Rule #1)
- Directions for naming:**
1. Name the more electronegative element first.
 - a. Oxygen is more electronegative than carbon so oxygen is named first.
 - b. The name of oxygen is changed to oxide.
 2. Name the less electronegative element as the second element.
 - a. The name of carbon is changed to carbide.
 - b. The number of atoms of the less electronegative element is named first.
 - c. The name of the element is changed to anion form (the name of the element with an -ide suffix).

Binary Compounds	
First Element	Second
1	mono
2	di
3	tri
4	tetra
5	penta
6	hexa
7	hepta
8	octa
9	nona
10	deca

Common Polyatomic Ions	
Formula	Name
<chem>OH-</chem>	hydroxide
<chem>NO3-</chem>	nitrate
<chem>SO4^2-</chem>	sulfate
<chem>CO3^2-</chem>	carbonate
<chem>PO4^3-</chem>	phosphate
<chem>NO2-</chem>	nitrite
<chem>SO3^2-</chem>	sulfite
<chem>PO3^3-</chem>	phosphite
<chem>NO-</chem>	nitrosyl
<chem>NO2+</chem>	nitrosonium

Naming Compounds with a Metal and a Nonmetal

- EXAMPLE**
- CaCl2
 Calcium Chloride
- Directions for naming:**
1. Name the metal as the first element. (No change)
 2. Name the nonmetal as the second element. (Change the ending to -ide)

Naming System			
Element	Number of Atoms	Element	Number of Atoms
Iron	1	Oxygen	2
Iron	2	Oxygen	1
Iron	3	Oxygen	1
Iron	4	Oxygen	1
Iron	5	Oxygen	1
Iron	6	Oxygen	1
Iron	7	Oxygen	1
Iron	8	Oxygen	1
Iron	9	Oxygen	1
Iron	10	Oxygen	1

Common Names for Specific Compounds			
Formula	Common Name	Formula	Common Name
<chem>NaCl</chem>	table salt	<chem>NaOH</chem>	caustic soda
<chem>KCl</chem>	salt	<chem>CaCl2</chem>	calcium chloride
<chem>NaNO3</chem>	potassium nitrate	<chem>CaCO3</chem>	limestone
<chem>Na2CO3</chem>	soda ash	<chem>CaSO4</chem>	gypsum

Chemistry Naming Compounds Study Guide

Leo J. Malone



Chemistry Naming Compounds Study Guide:

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Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999. Provides worked-out solutions to text problems along with chapter-by-chapter outlines and a variety of self-tests at the end of each chapter.

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learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999

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This book delves into Chemistry Naming Compounds Study Guide. Chemistry Naming Compounds Study Guide is a crucial topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemistry Naming Compounds Study Guide, encompassing both the fundamentals and more intricate discussions.

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- Chapter 5: Conclusion

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