

ANSWER KEY

Workbook SCH 3U – Unit 2: Quantities in Chemical Reactions

Mole Calculations:

Part 1: Molar Mass

Use the periodic table to find the molar masses of the following.



$$1.0 \text{ g/mol} + 35.5 \text{ g/mol} = 36.5 \text{ g/mol}$$



$$2(39.1 \text{ g/mol}) + 12.0 \text{ g/mol} + 3(16.0 \text{ g/mol}) = 138.2 \text{ g/mol}$$



$$40.1 \text{ g/mol} + 2(16.0 \text{ g/mol}) + 2(1.0 \text{ g/mol}) = 74.1 \text{ g/mol}$$



$$3(23.0 \text{ g/mol}) + 31.0 \text{ g/mol} + 4(16.0 \text{ g/mol}) = 164.0 \text{ g/mol}$$

Part 2: Mole Conversions

Work each of the following problems. SHOW ALL WORK.

1. How many atoms are in 6.2 moles of aluminum?

$$? \text{ atoms Al} = 6.2 \text{ mol Al} \times \frac{6.02 \times 10^{23} \text{ atoms Al}}{1 \text{ mol Al}} = 3.7 \times 10^{24} \text{ atoms Al}$$

2. Convert 5.3×10^{25} molecules of CO_2 to moles.

$$? \text{ moles CO}_2 = 5.3 \times 10^{25} \text{ molecules CO}_2 \times \frac{1 \text{ mole CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 88 \text{ mol CO}_2$$

3. How many formula units of sodium acetate are in 0.87 moles of sodium acetate?

$$? \text{ fu. NaC}_2\text{H}_3\text{O}_2 = 0.87 \text{ mol NaC}_2\text{H}_3\text{O}_2 \times \frac{6.02 \times 10^{23} \text{ fu. NaC}_2\text{H}_3\text{O}_2}{1 \text{ mol NaC}_2\text{H}_3\text{O}_2} = 5.2 \times 10^{23} \text{ fu. NaC}_2\text{H}_3\text{O}_2$$

4. Convert 3.55 moles NaCl to formula units.

$$? \text{ fu. NaCl} = 3.55 \text{ mol NaCl} \times \frac{6.02 \times 10^{23} \text{ fu. NaCl}}{1 \text{ mol NaCl}} = 2.14 \times 10^{24} \text{ fu. NaCl}$$

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