



Reaction Rates and Equilibrium Study Guide

ANSWER KEY

- Describe how each of the following factors would affect the rate of a chemical reaction in terms of "particle collisions" in your response.
 - Temperature decreases** *If temperature decreases, the particles will move slower and there will be less reactant particle collisions AND a smaller percentage of particles will have enough energy to produce a reaction. The reaction rate will therefore decrease.*
 - Addition of iodide ions, I^- , as a catalyst** *The presence of the catalyst provides an alternative pathway for the reaction taking place. This means that a larger percentage of particles will have enough energy to produce a reaction. The reaction rate will therefore increase.*
 - Increased reactant concentration** *With more reactant particles present, there are more collisions between those particles. This means the reaction rate will increase.*
 - Crushing a solid into a powder (with larger surface area)** *As a powder, there is more surface area available for collisions to take place. This means there will be more collisions and therefore a faster reaction rate.*
- The reaction progress diagram shows the energy changes that take place when nitrogen dioxide gas reacts with carbon monoxide gas to form carbon dioxide gas and nitrogen monoxide gas.

Chemistry Reaction Rates And Equilibrium Study Guide

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Chemistry Reaction Rates And Equilibrium Study Guide:

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