



CHEMISTRY

Class 12 - Chemistry

1. A reaction is first order in X and second order in Y. [2]
- How is the rate affected on increasing the concentration of Y three times?
 - How is the rate affected when the concentrations of both X and Y are doubled?

2. The rate Law for the reaction $A + B \rightarrow C$ is rate = $k [A]^2 [B]$. What would the reaction rate be when concentration of both A and B are doubled? [2]

3. Define order of reaction. Predict the order of reaction in the given graphs: [2]

a.



b.



where $[R]_0$ is the initial concentration of reactant and $t_{1/2}$ is half-life.

4. What is the effect of catalyst on activation energy? Why? [2]
5. a. For a reaction $R \rightarrow P$ half life ($t_{1/2}$) is independent of the initial concentration of reactants. What is the order of the reaction? [2]
- b. Write one difference between the order and molecularity of a reaction.
6. Derive an expression to calculate time required for completion of zero order reaction. [2]
7. What is meant by elementary step in a reaction? [2]
8. Derive integrated rate equation for rate constant of a first order reaction. [2]
9. Calculate the half life of first order reaction whose rate constant is 200 s^{-1} . [2]
10. A reaction is of second order with respect to a reactant. Now, is its rate affected if the concentration of reactant is: [2]
- doubled
 - reduced to half
11. The following data were obtained during the first order thermal decomposition of N_2O_5 (g) at constant volume: [3]
- $$2N_2O_5 (g) \rightarrow 2N_2O_4 (g) + O_2 (g)$$

Chemistry Reaction Rates Study Guide Answers

Léo Pomar



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Table of Contents Chemistry Reaction Rates Study Guide Answers

1. Understanding the eBook Chemistry Reaction Rates Study Guide Answers
 - The Rise of Digital Reading Chemistry Reaction Rates Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Reaction Rates Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Reaction Rates Study Guide Answers
 - Personalized Recommendations
 - Chemistry Reaction Rates Study Guide Answers User Reviews and Ratings

- Chemistry Reaction Rates Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Reaction Rates Study Guide Answers Free and Paid eBooks
 - Chemistry Reaction Rates Study Guide Answers Public Domain eBooks
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 - Highlighting and Note-Taking Chemistry Reaction Rates Study Guide Answers
 - Interactive Elements Chemistry Reaction Rates Study Guide Answers
- 8. Staying Engaged with Chemistry Reaction Rates Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Reaction Rates Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Reaction Rates Study Guide Answers
 - Setting Reading Goals Chemistry Reaction Rates Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Reaction Rates Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Reaction Rates Study Guide Answers
 - Distinguishing Credible Sources

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14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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